

# REPORT FOR EXPERIMENT 12

INSTRUCTOR

## Single Displacement Reactions

Evidence of Reaction	Equations (to be completed)
Describe any evidence of reaction; if no reaction was observed, write "None".	Write "No reaction", if no reaction was observed.
1. <u>Spangy, dark red/black</u>	$\text{Cu} + \text{AgNO}_3(\text{aq}) \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$
2. <u>Dark precipitate</u>	$\text{Pb} + \text{Cu(NO}_3)_2(\text{aq}) \rightarrow \text{Pb(NO}_3)_2 + \text{Cu}$
3. <u>Spangy, dark</u>	$\text{Zn} + \text{Pb(NO}_3)_2(\text{aq}) \rightarrow \text{Zn(NO}_3)_2 + \text{Pb}$
4. <u>none</u>	$\text{Zn} + \text{MgSO}_4(\text{aq}) \rightarrow \text{no reaction}$
5. <u>none</u>	$\text{Cu} + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{no reaction}$
6. <u>Same color spangy texture</u>	$\text{Zn} + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{ZnSO}_4 + \text{H}_2$

## QUESTIONS AND PROBLEMS

- Complete the following table by writing the symbols of the two elements whose reactivities are being compared in each test:

	Tube Number					
	1	2	3	4	5	6
Greater activity	Cu	Pb	Zn	Mg	H	Zn
Lesser activity	Ag	Cu	Pb	Zn	Cu	H

- Arrange Pb, Mg, and Zn in order of their activities, listing the most active first.

- (1) Mg
- (2) Zn
- (3) Pb

# Report For Experiment 12 Single Displacement Reactions

**K Payea**



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### **Table of Contents Report For Experiment 12 Single Displacement Reactions**

1. Understanding the eBook Report For Experiment 12 Single Displacement Reactions
  - The Rise of Digital Reading Report For Experiment 12 Single Displacement Reactions
  - Advantages of eBooks Over Traditional Books
2. Identifying Report For Experiment 12 Single Displacement Reactions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Report For Experiment 12 Single Displacement Reactions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Report For Experiment 12 Single Displacement Reactions

- Personalized Recommendations
- Report For Experiment 12 Single Displacement Reactions User Reviews and Ratings
- Report For Experiment 12 Single Displacement Reactions and Bestseller Lists
- 5. Accessing Report For Experiment 12 Single Displacement Reactions Free and Paid eBooks
  - Report For Experiment 12 Single Displacement Reactions Public Domain eBooks
  - Report For Experiment 12 Single Displacement Reactions eBook Subscription Services
  - Report For Experiment 12 Single Displacement Reactions Budget-Friendly Options
- 6. Navigating Report For Experiment 12 Single Displacement Reactions eBook Formats
  - ePub, PDF, MOBI, and More
  - Report For Experiment 12 Single Displacement Reactions Compatibility with Devices
  - Report For Experiment 12 Single Displacement Reactions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Report For Experiment 12 Single Displacement Reactions
  - Highlighting and Note-Taking Report For Experiment 12 Single Displacement Reactions
  - Interactive Elements Report For Experiment 12 Single Displacement Reactions
- 8. Staying Engaged with Report For Experiment 12 Single Displacement Reactions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Report For Experiment 12 Single Displacement Reactions
- 9. Balancing eBooks and Physical Books Report For Experiment 12 Single Displacement Reactions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Report For Experiment 12 Single Displacement Reactions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Report For Experiment 12 Single Displacement Reactions
  - Setting Reading Goals Report For Experiment 12 Single Displacement Reactions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Report For Experiment 12 Single Displacement Reactions

- Fact-Checking eBook Content of Report For Experiment 12 Single Displacement Reactions
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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