

CHAPTER 6 REVIEW*Chemical Bonding***SECTION 6-1****SHORT ANSWER** Answer the following questions in the space provided.

1. a A chemical bond between atoms results from the attraction between electrons and _____.
(a) protons (c) isotopes
(b) neutrons (d) Lewis structures
2. b A covalent bond consists of _____.
(a) a shared electron (c) two different ions
(b) a shared electron pair (d) an octet of electrons
3. a If two covalently bonded atoms are identical, the bond is identified as _____.
(a) nonpolar covalent (c) nonionic
(b) polar covalent (d) dipolar
4. b A covalent bond in which there is an unequal attraction for the shared electrons is _____.
(a) nonpolar (c) ionic
(b) polar (d) dipolar
5. c Atoms with a strong attraction for electrons they share with another atom exhibit _____.
(a) zero electronegativity (c) high electronegativity
(b) low electronegativity (d) Lewis electronegativity
6. c Bonds that possess between 5% and 50% ionic character are considered to be _____.
(a) ionic (c) polar covalent
(b) pure covalent (d) nonpolar covalent
7. a The greater the electronegativity difference between two atoms bonded together, the greater the bond's percentage of _____.
(a) ionic character (c) metallic character
(b) covalent character (d) electron sharing
8. The electrons involved in the formation of a chemical bond are called _____
valence electrons
9. A chemical bond that results from the electrostatic attraction between positive and negative ions is called a(n) _____
ionic bond

Prentice Hall Chemistry Review Bonds Answer Key

RJ Alexander



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