

Study Guide

Key Concepts

14.1 Properties of Gases

- Gases are easily compressed because of the space between particles in a gas.
- The amount of gas (n), volume (V), and temperature (T) are factors that affect gas pressure (P).

14.2 The Gas Laws

- As the pressure of a gas increases, the volume decreases, if the temperature is constant.
- As the temperature of an enclosed gas increases, the volume increases, if the pressure is constant.
- As the temperature of an enclosed gas increases, the pressure increases, if the volume is constant.

- The combined gas law allows you to do calculations for situations in which only the amount of gas is constant.

14.3 Ideal Gases

- To calculate the number of moles of a contained gas requires an expression that contains the variable n .
- Real gases differ most from an ideal gas at low temperatures and high pressures.

14.4 Gases: Mixtures and Movements

- In a mixture of gases, the total pressure is the sum of the partial pressures of the gases.
- Gases of lower molar mass diffuse and effuse faster than gases of higher molar mass.

Vocabulary

- Boyle's law (p. 418)
- Charles's law (p. 420)
- combined gas law (p. 424)
- compressibility (p. 413)
- Dalton's law of partial pressures (p. 432)
- diffusion (p. 435)
- effusion (p. 435)
- Gay-Lussac's law (p. 422)
- Graham's law of effusion (p. 435)
- ideal gas constant (p. 426)
- ideal gas law (p. 426)
- partial pressure (p. 432)

Key Equations

- Boyle's law: $P_1 \times V_1 = P_2 \times V_2$
- Charles's law: $\frac{V_1}{T_1} = \frac{V_2}{T_2}$
- Gay-Lussac's law: $\frac{P_1}{T_1} = \frac{P_2}{T_2}$
- Combined gas law: $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$
- Ideal gas law: $P \times V = n \times R \times T$ or $PV = nRT$
- Dalton's law: $P_{\text{total}} = P_1 + P_2 + P_3 + \dots$
- Graham's law: $\frac{\text{Rate}_A}{\text{Rate}_B} = \sqrt{\frac{\text{molar mass}_B}{\text{molar mass}_A}}$

Organizing Information

Use these terms to construct a concept map that organizes the major ideas of this chapter.



Concept Map 14 Solve the Concept Map with the help of an interactive guided tutorial.

with **ChemASAP**

Boyle's law

Charles's law

partial pressure

combined gas law

diffusion

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Table of Contents Prentice Hall Chemistry Chapter 14 Packet Answers

1. Understanding the eBook Prentice Hall Chemistry Chapter 14 Packet Answers
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 14 Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 14 Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 14 Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 14 Packet Answers
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 14 Packet Answers User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 14 Packet Answers and Bestseller Lists
5. Accessing Prentice Hall Chemistry Chapter 14 Packet Answers Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 14 Packet Answers Public Domain eBooks
 - Prentice Hall Chemistry Chapter 14 Packet Answers eBook Subscription Services
 - Prentice Hall Chemistry Chapter 14 Packet Answers Budget-Friendly Options

6. Navigating Prentice Hall Chemistry Chapter 14 Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 14 Packet Answers Compatibility with Devices
 - Prentice Hall Chemistry Chapter 14 Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 14 Packet Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 14 Packet Answers
 - Interactive Elements Prentice Hall Chemistry Chapter 14 Packet Answers
8. Staying Engaged with Prentice Hall Chemistry Chapter 14 Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 14 Packet Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 14 Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 14 Packet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 14 Packet Answers
 - Setting Reading Goals Prentice Hall Chemistry Chapter 14 Packet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 14 Packet Answers
 - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 14 Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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