

11-3 Practice Problems

1. Identify the limiting reactant when 1.22-g of O_2 reacts with 1.65-g of H_2 to produce water.

2. Identify the limiting reactant when 4.68-g of Fe reacts with 2.88-g of S to produce FeS .

3. Identify the limiting reactant when 2.87-g of Mg(OH)_2 reacts with 12.84-g of HCl to form MgCl_2 and water.

4. Identify the limiting reactant when 6.25-g of AgNO_3 reacts with 4.12-g of NaCl to form NaNO_3 and AgCl .

5. Identify the limiting reactant when 7.81-g of HCl reacts with 5.24-g of NaOH to produce NaCl and H_2O .

6. Identify the limiting reactant when 6.33-g of H_2SO_4 reacts with 5.92-g of NaOH to produce Na_2SO_4 and water.

7. Identify the limiting reactant when 49.25-g of CaC_2 reacts with 33.71-g of water to produce Ca(OH)_2 and C_2H_2 .

8. Identify the limiting reactant when 63.14-g of CaCl_2 reacts with 74.68-g of Na_2CO_3 to produce CaCO_3 and NaCl .

9. Identify the limiting reactant when 4.687-g of SF_4 reacts with 6.281-g of I_2O_5 to produce IF_3 and SO_2 .

10. If 4.1-g of Cr is heated with 9.3-g of Cl_2 , what mass CrCl_3 will be produced?

11. What mass of SO_2 is produced from the reaction between 31.5-g of S_8 and 8.65-g of O_2 ?

12. What mass of SO_2 is produced from the reaction of 12.4-g of SO_2 and 3.45-g of O_2 ?

13. What mass of H_2SO_4 is produced from the reaction of 6.58-g of SO_3 and 1.64-g of H_2O ?

14. What mass of CdS is produced if 8.47-g of cadmium reacts with 2.51-g of Sulfur?

Prentice Hall Chemistry Chapter 11 Test Answers

AW Rasmussen



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