

11

CHEMICAL REACTIONS

Practice Problems

In your notebook, solve the following problems. Use the 3-step problem-solving approach you learned in Chapter 1.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS

1. Write the skeleton equation for the reaction between hydrogen and oxygen that produces water.
2. Write the skeleton equation for the reaction that produces iron(II) sulfide from iron and sulfur.
3. Write the skeleton equation representing the heating of magnesium carbonate to produce solid magnesium oxide and carbon dioxide gas.
4. Write a balanced equation for the production of HCl gas from its elements.
5. Write a sentence that completely describes the chemical reaction represented by this balanced equation.



6. Write the word equation for the following equation. Write a sentence fully describing the reaction. Is the equation correctly balanced? Explain.



7. Write a balanced equation representing the formation of aqueous sulfuric acid from water and sulfur trioxide gas.
8. Write a balanced equation from this word equation.
aqueous silver nitrate + copper metal $\rightarrow$ silver metal + aqueous copper nitrate
9. Write a balanced equation for the following word equation.



SECTION 11.2 TYPES OF CHEMICAL REACTIONS

1. Write a balanced equation representing the reaction of magnesium with oxygen gas to produce magnesium oxide.
2. Write the balanced equation for the reaction that occurs between aluminum and fluorine.
3. Write the balanced equation for the production of oxygen gas and potassium chloride from the decomposition of potassium chlorate.
4. Write the balanced equation for the reaction between hydrochloric acid and calcium metal. The products are hydrogen gas and calcium chloride.
5. Write the balanced equation for the combustion of propane (C_3H_8) to produce carbon dioxide and water vapor.
6. Write the balanced equation for the reaction between iron(III) chloride and sodium hydroxide. The products are iron(III) hydroxide and sodium chloride.

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Puja Mehta

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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Prentice Hall Chemistry, 2000 Addison-Wesley Chemistry Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, Prentice-Hall Staff, 2001-02 To purchase or download a workbook click on the Purchase or Download button to the left To purchase a workbook enter the desired quantity and click Add to Cart To download a free workbook right click the FREE Download PDF link and save to your computer This will result in a faster download as opposed to left clicking and opening the link **Water Chemistry** Stanley E. Manahan, 2010-08-19 Carefully crafted to provide a comprehensive overview of the chemistry of water in the environment Water Chemistry Green Science and Technology of Nature s Most Renewable Resource examines water issues within the broad framework of sustainability an issue of increasing importance as the demands of Earth s human population threaten to overwhelm the planet s carrying capacity Renowned environmental author Stanley Manahan provides more than just basic coverage of the chemistry of water He relates the science and technology of this amazing substance to areas essential to sustainability science including environmental and green chemistry industrial ecology and green sustainable science and technology The inclusion of a separate chapter that comprehensively covers energy including renewable and emerging sources sets this book a part Manahan explains how the hydrosphere relates to the geosphere atmosphere biosphere and anthrosphere His approach views Planet Earth as consisting of these five mutually interacting spheres He covers biogeochemical cycles and the essential role of water in these basic cycles of materials He also defines environmental chemistry and green chemistry emphasizing water s role in the practice of each Manahan highlights the role of the anthrosphere that part of the environment constructed and operated by humans He underscores its overwhelming influence on the environment and its pervasive effects on the hydrosphere He also covers the essential role that water plays in the sustainable operation of the anthrosphere and how it can be maintained in a manner that will enable it to operate in harmony with the environment for generations to come Written at an intermediate level this is an appropriate text for the study of current affairs in environmental chemistry It provides a review and grounding in basic and organic chemistry for those students who need it and also fills a niche for an aquatic chemistry book that relates the hydrosphere to the four other environmental spheres **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Philosophy of Science for Nursing Practice** Michael D. Dahnke, Heyward Michael Dreher, 2011 Print CourseSmart **Hydrometallurgy** Michael L. Free, 2021-11-30 This revised new edition retains its class tested coverage of how metals behave in water while updating and expanding information about

metals processing methods The book further retains its emphasis on predicting and engineering the way metals are extracted from ore sources separated from unwanted entities recovered as metals and purified using water based processing The transformation of minerals to metals requires hydrometallurgical processing for nearly all of the nonferrous metals we use This book elucidates the associated fundamentals and processing applications as well as related tools to assess processes and performance The new edition further includes additional photographs updated drawings supplementary data updated descriptive information and new detail on rare earth elements processing as well as recycling and byproduct recovery of metals

Introductory Chemistry Charles H. Corwin, 2011 To help students learn chemical skills and concepts more effectively Introductory Chemistry Concepts and Critical Thinking Sixth Edition highlights the connection between key concepts and key problem solving skills through critical thinking Math and problem solving are covered early in the text Corwin builds your problem solving ability through innovative learning aids and technology formulated to meet your needs This revision retains all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking

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Problems and Problem Solving in Chemistry Education Georgios Tsapalis, 2021 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools

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