

11-2 Practice Problems

1. Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:



$$\frac{0.38\text{g Li}_3\text{N}}{1} \times \frac{1\text{mol Li}_3\text{N}}{35\text{g Li}_3\text{N}} \times \frac{3\text{mol LiOH}}{1\text{mol Li}_3\text{N}} \times \frac{24\text{g LiOH}}{1\text{mol LiOH}} = \boxed{0.78\text{g LiOH}}$$

2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?



$$\frac{0.29\text{g NaI}}{1} \times \frac{1\text{mol}}{150} \times \frac{2\text{mol NaCl}}{2\text{mol NaI}} \times \frac{58\text{g}}{1\text{mol}} = \boxed{0.112\text{g NaCl}}$$

3. Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:



$$\frac{0.85\text{g C}_4\text{H}_{10}}{1} \times \frac{1\text{mol}}{58\text{g}} \times \frac{8\text{mol CO}_2}{2\text{mol C}_4\text{H}_{10}} \times \frac{44\text{g}}{1\text{mol}} = \boxed{2.58\text{g CO}_2}$$

4. Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:



$$\frac{0.46\text{g Sb}_2\text{O}_3}{1} \times \frac{1\text{mol}}{291.52\text{g}} \times \frac{2\text{mol Sb}}{1\text{mol Sb}_2\text{O}_3} \times \frac{121.76\text{g}}{1\text{mol}} = \boxed{0.38\text{g Sb}}$$

5. What mass of hydrogen peroxide (H_2O_2) must decompose to produce 0.77 g of water?



$$\frac{0.77\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{2\text{mol H}_2\text{O}_2}{2\text{mol H}_2\text{O}} \times \frac{34\text{g}}{1\text{mol}} = \boxed{1.45\text{g H}_2\text{O}_2}$$

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?



$$\frac{0.69\text{g CO}_2}{1} \times \frac{1\text{mol}}{44\text{g}} \times \frac{2\text{mol CO}}{2\text{mol CO}_2} \times \frac{28\text{g}}{1\text{mol}} = \boxed{0.44\text{g CO}}$$

7. Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:



$$\frac{0.73\text{g Ni}(\text{NO}_3)_2}{1} \times \frac{1\text{mol}}{185\text{g}} \times \frac{2\text{mol NaNO}_3}{1\text{mol Ni}(\text{NO}_3)_2} \times \frac{85\text{g}}{1\text{mol}} = \boxed{0.68\text{g NaNO}_3}$$

8. Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:



$$\frac{0.64\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{1\text{mol Ca}(\text{OH})_2}{2\text{mol H}_2\text{O}} \times \frac{74\text{g}}{1\text{mol}} = \boxed{1.32\text{g Ca}(\text{OH})_2}$$

9. How many grams of ozone (O_3) must decompose to produce 0.87 g of oxygen?



$$\frac{0.87\text{g O}_2}{1} \times \frac{1\text{mol}}{32\text{g O}_2} \times \frac{2\text{mol}}{3\text{mol}} \times \frac{48\text{g O}_3}{1\text{mol}} = \boxed{0.87\text{g O}_3}$$

10. Find the mass of sugar ($\text{C}_6\text{H}_{12}\text{O}_6$) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation:



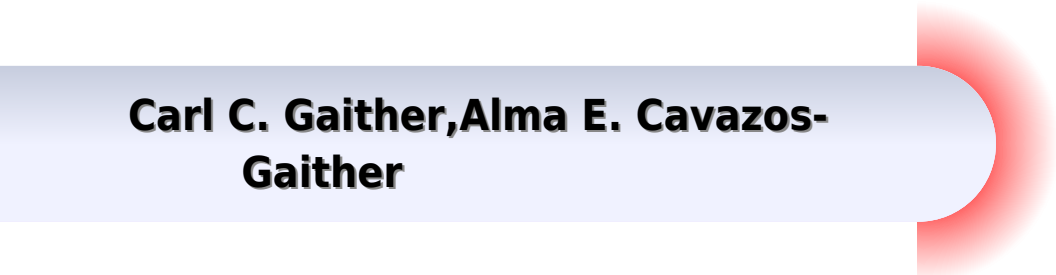
11. How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$.

12. Find the mass of benzene (C_6H_6) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:



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**Carl C. Gaither, Alma E. Cavazos-
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Engineering Phillip C. Wankat, 2022-10-24 The Definitive Learner Friendly Guide to Chemical Engineering Separations Extensively Updated Including a New Chapter on Melt Crystallization Efficient separation processes are crucial to addressing many societal problems from developing new medicines to improving energy efficiency and reducing emissions Separation Process Engineering Fifth Edition is the most comprehensive accessible guide to modern separation processes and the fundamentals of mass transfer In this completely updated edition Phillip C Wankat teaches each key concept through detailed realistic examples using actual data with up to date simulation practice spreadsheet based exercises and references Wankat thoroughly covers each separation process including flash column and batch distillation exact calculations and shortcut methods for multicomponent distillation staged and packed column design absorption stripping and more His extensive discussions of mass transfer and diffusion enable faculty to teach separations and mass transfer in a single course And detailed material on liquid liquid extraction adsorption chromatography and ion exchange prepares students for advanced work New and updated content includes melt crystallization steam distillation residue curve analysis batch washing the Shanks system for percolation leaching eutectic systems forward osmosis microfiltration and hybrid separations A full chapter discusses economics and energy conservation including updated equipment costs Over 300 new and updated homework problems are presented all extensively tested in undergraduate courses at Purdue University New chapter on melt crystallization solid liquid phase equilibrium suspension static and falling film layer approaches and 34 questions and problems New binary VLE equations and updated content on simultaneous solutions New coverage of safety and fire hazards New material on steam distillation simple multi component batch distillation and residue curve analysis Expanded discussion of tray efficiencies packed column design and energy reduction in distillation New coverage of two hybrid extraction with distillation and the Kremser equation in fractional extraction Added sections on deicing with eutectic systems eutectic freeze concentration and scale up New sections on forward osmosis and microfiltration Expanded advanced content on adsorption and ion exchange including updated instructions for eight detailed Aspen Chromatography labs Discussion of membrane separations including gas permeation reverse osmosis ultrafiltration pervaporation and applications Thirteen up to date Aspen Plus process simulation labs adaptable to any simulator This guide reflects an up to date understanding of how

modern students learn designed organized and written to be exceptionally clear and easy to use It presents detailed examples in a clear standard format using real data to solve actual engineering problems preparing students for their future careers

Computational Chemistry Errol G. Lewars, 2016-09-20 This is the third edition of the successful text reference book that covers computational chemistry It features changes to the presentation of key concepts and includes revised and new material with several expanded exercises at various levels such as harder questions for those ready to be tested in greater depth this aspect is absent from other textbooks in the field Although introductory and assuming no prior knowledge of computational chemistry it covers the essential aspects of the subject There are several introductory textbooks on computational chemistry this one is as in its previous editions a unique textbook in the field with copious exercises and questions and solutions with discussions Noteworthy is the fact that it is the only book at the introductory level that shows in detail yet clearly how matrices are used in one important aspect of computational chemistry It also serves as an essential guide for researchers and as a reference book

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School Life ,1962

Elements of Chemistry Robert S.

Boikess, Kenneth Breslauer, Edward Edelson, 1986 Voyage of Adventure. Annotated Teacher's Edition ,1986 **Chemical Problem Solving Using Dimensional Analysis** Robert Nakon, 1988 **Geochemistry** Harry Y. McSween, Steven McAfee Richardson, Maria E. Uhle, 2003 Written expressly for undergraduate and graduate geologists this book focuses on how geochemical principles can be used to solve practical problems The attention to problem solving reflects the authors belief that showing how theory is useful in solving real life problems is vital for learning The book gives students a thorough grasp of the basic principles of the subject balancing the traditional equilibrium perspective and the kinetic viewpoint The first half of the book considers processes in which temperature and pressure are nearly constant After introductions to the laws of thermodynamics to fundamental equations for flow and diffusion and to solution chemistry these principles are used to investigate diagenesis weathering and natural waters The second half of the book applies thermodynamics and kinetics to systems undergoing changes in temperature and pressure during magmatism and metamorphism This revised edition incorporates new geochemical discoveries as examples of processes and pathways with new chapters on mineral structure and bonding and on organic matter and biomarkers Each chapter has worked problems and the authors assume that the student has had a year of college level chemistry and a year of calculus Praise for the first edition A truly modern geochemistry book Very well written and quite enjoyable to read An excellent basic text for graduate level instruction in geochemistry Journal of Geological Education An up to date broadly conceived introduction to geochemistry Given the recent flowering of geochemistry as an interdisciplinary science and given the extent to which it now draws upon the fundamentals of thermodynamics and kinetics to understand earth and planetary processes this timely and rigorous book is welcome indeed *Geochimica et Cosmochimica Acta* *Clear and Simple Chemistry* James Berger, Jim Berger, 1986 Contemporary Chemical Analysis Judith F. Robinson, Kenneth A. Robinson, 1998 This book provides a readable yet rigorous introduction to analytical methods with a focus on problem solving skills It stresses the fundamental concepts of chemical analysis and through examples from current journals and other science media shows how the principles and practice of analytical chemistry are used to produce answers to questions in all areas of scientific study and practice Features a balance of topics that is closer to contemporary analytical practice than those covered by other books Introduces the tools that are ubiquitous in analytical chemistry e g statistics sampling and sample preparation Discusses methods depending on chemical kinetics which are so widely used in medicine and biology Features a number of problems that call for the use of a spreadsheet to generate data which is then plotted to show trends Includes answers for all numerical problems in an appendix

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