

Name \_\_\_\_\_

## 7-1 Review and Reinforcement

### *Ionic Bonding*

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative  
positive  
octet rule

noble gas  
electron  
atom

octet  
ions  
cation

electron configuration  
polyatomic  
empirical formula

1. An ionic bond is an attraction between oppositely charged \_\_\_\_\_.
2. Anions have a \_\_\_\_\_ charge.
3. An atom becomes an ion by losing or gaining \_\_\_\_\_.
4. The \_\_\_\_\_ rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) \_\_\_\_\_.
6. The atoms that make up \_\_\_\_\_ ions are bonded together by covalent bonds.
7. The \_\_\_\_\_ method can be used to write the formula for an ionic compound.
8. The \_\_\_\_\_ of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- |       |                                                                                              |
|-------|----------------------------------------------------------------------------------------------|
| _____ | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons.                       |
| _____ | 10. A cation has a <u>positive</u> charge.                                                   |
| _____ | 11. A binary ionic compound contains only <u>one</u> kind of cations and one kind of anions. |
| _____ | 12. <u>Monatomic</u> ions consist of more than one atom.                                     |
| _____ | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons.        |
| _____ | 14. A great deal of energy is <u>produced</u> when an ionic compound is formed.              |

# Prentice Hall 12 3 Review And Reinforcement Answer

**Lei Huang**



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