Name

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Date

Chapter 7 Chemical Reactions

Section 7.1 Describing Reactions

This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.

Reading Strategy (page 192)

Monitoring Your Understanding Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.

What I Expect to Learn	What I Learned
Artsetrs may vary. Possible answers: How to balance chemical equations	Answers may very. Possible answers: An unbelanced equation can be halanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the motar mass as a conversion factor.

Chemical Equations (pages 192-193)

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products.
- Circle the letter of each sentence that is a correct interpretation of the chemical equation C + O₃ → CO₃.
 - Carbon and oxygen react and form carbon monoxide.
 - (b) Carbon and oxygen react and form carbon dioxide.
 - c. Carbon dioxide yields carbon and oxygen.
 - (d) The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction.
- 4. Circle the letter of the correct answer. According to the equation C + O₂ → CO₂, how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
 - a. 1 (c)14

- b. 7 d. 28
- In the reaction represented by the equation C + O₂ → CO₂, the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted

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Questions 50 MCQs Group IV Multiple Choice Questions 53 MCQs Groups II and VII Multiple Choice Questions 180 MCQs Halogenoalkanes Multiple Choice Questions 33 MCQs Hydrocarbons Multiple Choice Questions 53 MCQs Introduction to Organic Chemistry Multiple Choice Questions 52 MCQs Ionic Equilibria Multiple Choice Questions 56 MCQs Lattice Energy Multiple Choice Questions 33 MCQs Moles and Equations Multiple Choice Questions 50 MCQs Nitrogen and Sulfur Multiple Choice Questions 89 MCQs Organic and Nitrogen Compounds Multiple Choice Questions 54 MCQs Periodicity Multiple Choice Questions 202 MCQs Polymerization Multiple Choice Questions 36 MCQs Rates of Reaction Multiple Choice Questions 39 MCQs Reaction Kinetics Multiple Choice Questions 52 MCQs Redox Reactions and Electrolysis Multiple Choice Questions 55 MCQs States of Matter Multiple Choice Questions 66 MCQs Transition Elements Multiple Choice Questions 29 MCQs The chapter Alcohols and Esters MCQs covers topics of introduction to alcohols and alcohols reactions The chapter Atomic Structure and Theory MCQs covers topics of atom facts elements and atoms number of nucleons protons electrons and neutrons The chapter Benzene Chemical Compound MCQs covers topics of benzene arenes reaction phenol properties and reactions of phenol The chapter Carbonyl Compounds MCQs covers topics of carbonyl compounds aldehydes and ketone testing nucleophilic addition with HCN preparation of aldehydes and ketone reduction of aldehydes and ketone **Prentice** Hall Science Explorer Chemical Interactions Adapted Reading and Study Workbook Michael J. Padilla, Ioannis Miaoulis, Martha Cyr, Prentice Hall Staff, 2005-05-01 1 Atoms and Bonding 2 Chemical Reactions 3 Acids Bases and Solutions 4 Carbon Chemistry

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