

11-3 Practice Problems

1. Identify the limiting reactant when 1.22-g of O_2 reacts with 1.65-g of H_2 to produce water.

2. Identify the limiting reactant when 4.68-g of Fe reacts with 2.88-g of S to produce FeS .

3. Identify the limiting reactant when 2.87-g of $\text{Mg}(\text{OH})_2$ reacts with 12.84-g of HCl to form MgCl_2 and water.

4. Identify the limiting reactant when 6.25-g of AgNO_3 reacts with 4.12-g of NaCl to form NaNO_3 and AgCl .

5. Identify the limiting reactant when 7.81-g of HCl reacts with 5.24-g of NaOH to produce NaCl and H_2O .

6. Identify the limiting reactant when 6.33-g of H_2SO_4 reacts with 5.92-g of NaOH to produce Na_2SO_4 and water.

7. Identify the limiting reactant when 49.25-g of CaC_2 reacts with 33.71-g of water to produce $\text{Ca}(\text{OH})_2$ and C_2H_2 .

8. Identify the limiting reactant when 63.14-g of CaCl_2 reacts with 74.68-g of Na_2CO_3 to produce CaCO_3 and NaCl .

9. Identify the limiting reactant when 4.687-g of SF_4 reacts with 6.281-g of I_2O_5 to produce IF_5 and SO_2 .

10. If 4.1-g of Cr is heated with 9.3-g of Cl_2 , what mass CrCl_3 will be produced?

11. What mass of SO_2 is produced from the reaction between 31.5-g of S_8 and 8.65-g of O_2 ?

12. What mass of SO_2 is produced from the reaction of 12.4-g of SO_2 and 3.45-g of O_2 ?

13. What mass of H_2SO_4 is produced from the reaction of 6.58-g of SO_3 and 1.64-g of H_2O ?

14. What mass of CdS is produced if 8.47-g of cadmium reacts with 2.51-g of Sulfur?

Prentice Hall Chemistry Practice Problem

M Tight



Prentice Hall Chemistry Practice Problem:

The book delves into Prentice Hall Chemistry Practice Problem. Prentice Hall Chemistry Practice Problem is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Prentice Hall Chemistry Practice Problem, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Prentice Hall Chemistry Practice Problem
- Chapter 2: Essential Elements of Prentice Hall Chemistry Practice Problem
- Chapter 3: Prentice Hall Chemistry Practice Problem in Everyday Life
- Chapter 4: Prentice Hall Chemistry Practice Problem in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Prentice Hall Chemistry Practice Problem. This chapter will explore what Prentice Hall Chemistry Practice Problem is, why Prentice Hall Chemistry Practice Problem is vital, and how to effectively learn about Prentice Hall Chemistry Practice Problem.

3. In chapter 2, the author will delve into the foundational concepts of Prentice Hall Chemistry Practice Problem. This chapter will elucidate the essential principles that need to be understood to grasp Prentice Hall Chemistry Practice Problem in its entirety.

4. In chapter 3, this book will examine the practical applications of Prentice Hall Chemistry Practice Problem in daily life. This chapter will showcase real-world examples of how Prentice Hall Chemistry Practice Problem can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Prentice Hall Chemistry Practice Problem in specific contexts. The fourth chapter will explore how Prentice Hall Chemistry Practice Problem is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about Prentice Hall Chemistry Practice Problem. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Prentice Hall Chemistry Practice Problem.

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