

REVIEW: Chapter 8 - Covalent Bonding

- Typically, how many bonds & lone pairs will each have when covalently bonding?
 - Boron – 3 bonds; 0 lone pairs
 - Silicone – 4 bonds; 0 lone pairs
 - Neon – 0 bonds; 4 lone pairs
 - Iodine – 1 bond; 3 lone pairs
 - Nitrogen – 3 bonds; 1 lone pair
 - Oxygen – 2 bonds; 2 lone pairs
- Which elements do not follow the octet rule? (Hint – there are four we went over)
 - H & B never follow the octet rule.
 - P & S sometimes don't follow the octet rule.
- What type of bond will each combination of atoms form?
 - F & N – 2 nonmetals, therefore covalent
 - Al & O – 1 metal/1 nonm., therefore ionic
 - Na & S – 1 metal/1 nonm., therefore ionic
 - B & Cl – 2 nonmetals, therefore covalent
 - Ca & I – 1 metal/1 nonm., therefore ionic
 - Fe & Cu – 2 metals, therefore metallic
- What is the difference between:
 - Molecules & Compounds?
Compounds can be made using all types of bonds (ionic, covalent or metallic), but molecules can only be made using covalent bonds
 - Monatomic elements & Diatomic molecules?
Monatomic elements exist as single atoms; diatomic molecules are made up of two of the same atom
 - Structural formulas, Empirical formulas, & Molecular formulas?
All three are for covalent bonds only! Molecular show the actual number of atoms in the molecule; Empirical give the lowest whole number ratio; Structural formulas are not really formulas, they are drawing of the molecule.
 - Polar bonds & Non-polar bonds?
Both describe covalent bonds! In a polar bond the electrons are NOT shared equally; in a non-polar bond electrons are shared equally
- Classify each of the following as an atom or a molecule:
 - Be, atom
 - CO, molec.
 - N₂, molec.
 - H₂O, molec.
 - Ne, atom

Prentice Hall Chemistry Chapter 8 Covalent Bonding

Answers

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Prentice Hall Chemistry Chapter 8 Covalent Bonding Answers:

Unveiling the Magic of Words: A Report on "**Prentice Hall Chemistry Chapter 8 Covalent Bonding Answers**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Prentice Hall Chemistry Chapter 8 Covalent Bonding Answers**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound affect on the souls of its readers.

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