

11-1 Practice Problems

1. Lead will react with Hydrochloric acid to produce Lead (II) Chloride and Hydrogen gas. How many moles of Hydrochloric acid are needed to completely react with 0.36-mol of lead?

2. How many moles of HNO_3 will be produced when 0.51-mol of N_2O_5 reacts according to the following equation?
 $\text{N}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3$

3. Iron will react with Oxygen Gas to produce Fe_2O_3 . How many moles of Fe_2O_3 will be produced if 0.18-mol of Fe reacts?

4. How many moles of NaBr will be produced when 0.69-mol of Bromine reacts according to the following equation?
 $\text{Br}_2 + 2\text{NaI} \rightarrow 2\text{NaBr} + \text{I}_2$

5. Phosphorus will react with Bromine to produce Phosphorus Tribromide. How many moles of Phosphorus Tribromide will be produced if 0.78-mol of Bromine is reacted?

6. Carbon will react with Zinc Oxide to produce Zinc and Carbon Dioxide. How many moles of carbon dioxide will be produced if 0.38-mol of ZnO is completely reacted?

7. How many moles of oxygen will be needed to react with 0.38-mol of C_3H_8 according to the following equation?
 $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

8. Nitrogen gas can react with Hydrogen gas to produce ammonia. How many moles of Nitrogen will be needed to produce 0.48-mol of NH_3 ?

9. How many moles of hydrogen will be produced if 0.44-mol of CaH_2 reacts according to the following equation?
 $\text{CaH}_2 + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + 2\text{H}_2$

10. How many moles of water will be produced if 2.35-mol of oxygen reacts according to the following equation?
 $2\text{C}_6\text{H}_6 + 15\text{O}_2 \rightarrow 12\text{CO}_2 + 6\text{H}_2\text{O}$

Prentice Hall Chemistry 11 2 Practice Problems Answers

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Prentice Hall Chemistry 11 2 Practice Problems Answers:

The book delves into Prentice Hall Chemistry 11 2 Practice Problems Answers. Prentice Hall Chemistry 11 2 Practice Problems Answers is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Prentice Hall Chemistry 11 2 Practice Problems Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Prentice Hall Chemistry 11 2 Practice Problems Answers
 - Chapter 2: Essential Elements of Prentice Hall Chemistry 11 2 Practice Problems Answers
 - Chapter 3: Prentice Hall Chemistry 11 2 Practice Problems Answers in Everyday Life
 - Chapter 4: Prentice Hall Chemistry 11 2 Practice Problems Answers in Specific Contexts
 - Chapter 5: Conclusion
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 6. In chapter 5, the author will draw a conclusion about Prentice Hall Chemistry 11 2 Practice Problems Answers. This chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Prentice Hall Chemistry 11 2 Practice Problems Answers.

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