

**Chapter 2 The Chemistry of Life****Section 2-1 The Nature of Matter (pages 35–39)****Key Concepts**

- What three subatomic particles make up atoms?
- How are all of the isotopes of an element similar?
- What are the two main types of chemical bonds?

**Atoms (page 35)**

1. The basic unit of matter is called a(an) \_\_\_\_\_.
2. Describe the nucleus of an atom. \_\_\_\_\_  
\_\_\_\_\_
3. Complete the table about subatomic particles.

**SUBATOMIC PARTICLES**

| Particle | Charge   | Location in Atom |
|----------|----------|------------------|
|          | Positive |                  |
|          | Neutral  |                  |
|          | Negative |                  |

4. Why are atoms neutral despite having charged particles? \_\_\_\_\_  
\_\_\_\_\_

**Elements and Isotopes (page 36)**

5. What is a chemical element? \_\_\_\_\_  
\_\_\_\_\_
6. What does an element's atomic number represent? \_\_\_\_\_  
\_\_\_\_\_
7. Atoms of the same element that differ in the number of neutrons they contain are known as \_\_\_\_\_.
8. How are isotopes identified? \_\_\_\_\_
9. Why do all isotopes of an element have the same chemical properties? \_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_

# Prentice Hall Biology Worksheets Chapter 9

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