

Predicting Products of Chemical Reactions

This worksheet is designed to help you predict products of simple reactions of the four basic reaction types (synthesis, decomposition, single replacement, and double replacement) and combustion reactions.

For the first few reactions, the type of reaction is listed, you should predict the products, then balance. Further questions just have the reactants listed and you should decide on the type of reaction, as well as the correct products. Many of these reactions fall into the category of redox reactions, though do not let that confuse you...each can be described in terms of the four basic reaction types (except the combustion reactions).

Although states (s, l, g, aq) of the reactants and products are very important in a chemical reaction, don't worry about determining those for these problems. Rather, focus on what products might result from the reactants given. Pay particular attention to the ionic charge of species that you know forms ions with only one possible charge (e.g., alkali metals, alkaline earth metals, halogens, etc.)

1. Combustion: $\text{C}_2\text{H}_6 + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
2. Combustion: $2 \text{C}_4\text{H}_{10} + 11 \text{O}_2 \rightarrow 8 \text{CO}_2 + 6 \text{H}_2\text{O}$
3. Combustion: $\text{C}_5\text{H}_{12}\text{O}_2 + 7 \text{O}_2 \rightarrow 6 \text{CO}_2 + 5 \text{H}_2\text{O}$
4. Synthesis: $\text{Mg} + \text{I}_2 \rightarrow \text{MgI}_2$
5. Double displacement: $\text{CuCl}_2 + \text{H}_2\text{S} \rightarrow \text{CuS} + 2 \text{HCl}$
6. Double displacement: $\text{NaOH} + \text{HClO}_4 \rightarrow \text{NaClO}_4 + \text{H}_2\text{O}$
7. Decomposition: $\text{ZnCO}_3 + \text{heat} \rightarrow \text{ZnO} + \text{CO}_2$
8. Single replacement: $2 \text{HCl} + \text{Zn} \rightarrow \text{ZnCl}_2 + \text{H}_2$
9. single $2 \text{Na} + \text{MgCl}_2 \rightarrow 2 \text{NaCl} + \text{Mg}$
10. double $\text{CaCl}_2 + \text{K}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + 2 \text{KCl}$
11. synthesis $2 \text{K} + \text{Cl}_2 \rightarrow 2 \text{KCl}$
12. double $3 \text{BaCl}_2 + 2 \text{K}_3\text{PO}_4 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6 \text{KCl}$
13. double $\text{H}_2\text{SO}_4 + 2 \text{KOH} \rightarrow \text{K}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
14. synthesis $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
15. double $\text{Pb}(\text{NO}_3)_2 + 2 \text{KOH} \rightarrow \text{Pb}(\text{OH})_2 + 2 \text{KNO}_3$
16. double $\text{H}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow 2 \text{HCl} + \text{BaSO}_4$
17. single $\text{Ca} + 2 \text{AgCl} \rightarrow \text{CaCl}_2 + 2 \text{Ag}$
18. double $\text{H}_3\text{PO}_4 + \text{FeBr}_3 \rightarrow 3 \text{HBr} + \text{FePO}_4$
19. synthesis $6 \text{Li} + \text{N}_2 \rightarrow 2 \text{Li}_3\text{N}$
20. double $2 \text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2 + 2 \text{H}_2\text{O}$

Predict Products And Balancing Equations Answers

ML Morrison



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Mastering NEET MCQs: Speed-solving Techniques for Medical Aspirants Dr. Ashokan Kannarath,2024-01-31 This book dives deep into the strategies techniques and tips required to excel in the NEET MCQs It is designed to help you not only solve questions quickly but also enhance your understanding and retention of vital medical concepts It explains how to analyse NEET MCQs without getting overwhelmed efficiently identifying keywords understanding question patterns and eliminating wrong answer options It also exposes proven strategies to increase your answering speed without compromising accuracy Includes time management tips mental math tricks and effective test taking methods

Engineering Chemistry : For the Students of Sethu Institute of Technology (SIT), Virudhunagar S S Dara & S S Umare, This book has been written as per the syllabus prescribed by Sethu Institute of Technology SIT Virudhunagar for the First Semester of Engineering Chemistry students The book has been developed in view of the recent development of the subject The book covers important topics such as Ionic and Electrovalent Bond Covalent Bond Variable Valency Coordinate or Dative Bond Complex Ions Chemical Equation Chemical Reactions Mathematical Representation Concept of pH Scale Rate of Reaction or Reaction Velocity Factors Influencing the Reaction Rate Rate Law or Rate Equation and Rate Constant Measurement of Rate of Reaction Order of a Reaction Pseudo Order Reactions Methods for Determination of Order of a Reaction Effect of Water on Rocks and Minerals Types and Effects of Impurities Present in Water Methods of Treatment of Water for Domestic Industrial Purpose Nernst Theory Standard Electrode Potentials Galvanic Series Reversible Cells Polarization How to Prevent Corrosion Electroplating etc have been explained in lucid manner The book is sincerely offered to students and teaching fraternities associated with engineering chemistry from various engineering and technological institutions all over the country

Sat Subject Test - Chemistry Kevin R. Reel,2006-07-10 Contains SAT test preparation for the subject of chemistry and offers six full length practice tests along with a chemistry review on all topics Also provides two full length practice tests on a CD ROM which are time and scored and includes explanations

Chemistry ,2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be

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