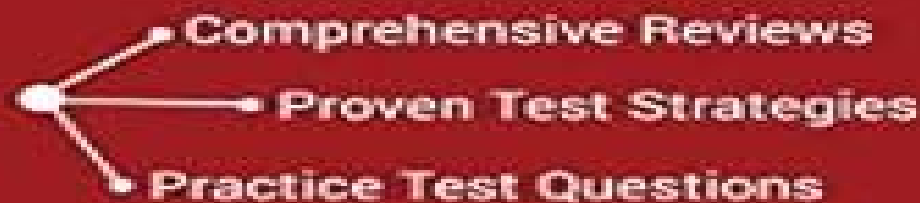


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Table of Contents Praxis 2 English Study Guide

1. Understanding the eBook Praxis 2 English Study Guide
 - The Rise of Digital Reading Praxis 2 English Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Praxis 2 English Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Praxis 2 English Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Praxis 2 English Study Guide
 - Personalized Recommendations
 - Praxis 2 English Study Guide User Reviews and Ratings
 - Praxis 2 English Study Guide and Bestseller Lists
5. Accessing Praxis 2 English Study Guide Free and Paid eBooks
 - Praxis 2 English Study Guide Public Domain eBooks

- Praxis 2 English Study Guide eBook Subscription Services
- Praxis 2 English Study Guide Budget-Friendly Options
- 6. Navigating Praxis 2 English Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Praxis 2 English Study Guide Compatibility with Devices
 - Praxis 2 English Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Praxis 2 English Study Guide
 - Highlighting and Note-Taking Praxis 2 English Study Guide
 - Interactive Elements Praxis 2 English Study Guide
- 8. Staying Engaged with Praxis 2 English Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Praxis 2 English Study Guide
- 9. Balancing eBooks and Physical Books Praxis 2 English Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Praxis 2 English Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Praxis 2 English Study Guide
 - Setting Reading Goals Praxis 2 English Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Praxis 2 English Study Guide
 - Fact-Checking eBook Content of Praxis 2 English Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho/para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no2 which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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