

Chemistry 1220
Practice Exam 3
7 July, 2003

Instructions:

Do not begin until 8:40 AM.

This exam should have 25 questions. Each question is worth 4 points for a total of 100 points. A periodic table should be on the reverse side of this page.

Use only a #2 pencil to mark the best answer on the scantron Sheet. You will turn in only the scantron sheet and your homework set 2. You should keep the rest of the exam.

On the scantron sheet, fill out your name and SS# and be sure to fill out the corresponding bubbles.

Information:

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ (at } 25^\circ\text{C)}$$

$$K_a \times K_b = K_w$$

$$\text{pH} = \text{p}K_a + \log ([\text{A}^-]/[\text{HA}])$$

$$\text{pH} + \text{pOH} = 14$$

$$\Delta S_{\text{sys}} = q_{\text{rev}}/T \text{ (at constant } T)$$

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta G^\circ + RT \ln Q$$

$$\Delta G^\circ = -RT \ln K_{\text{eq}}$$

$$1\text{V} = 1\text{J/C}$$

$$1\text{W} = 1\text{J/s}$$

$$\Delta G = -nFE$$

$$\Delta G^\circ = -nFE^\circ$$

$$1\text{F} = 96500\text{ C/mol} = 96,500\text{ J/(V} \times \text{mol)}$$

$$E^\circ_{\text{cell}} = E^\circ_{\text{red(cathode)}} - E^\circ_{\text{red(anode)}}$$

$$E^\circ_{\text{cell}} = E^\circ_{\text{red(cathode)}} + E^\circ_{\text{ox(anode)}}$$

$$E = E^\circ - (0.0592/n) \log Q$$

$$\log K_{\text{eq}} = (nE^\circ)/0.0592$$

$$w = nFE_{\text{ext}}$$

$$R = 8.314\text{ J/(mol} \times \text{K)}$$

$$K = ^\circ\text{C} + 273.15$$

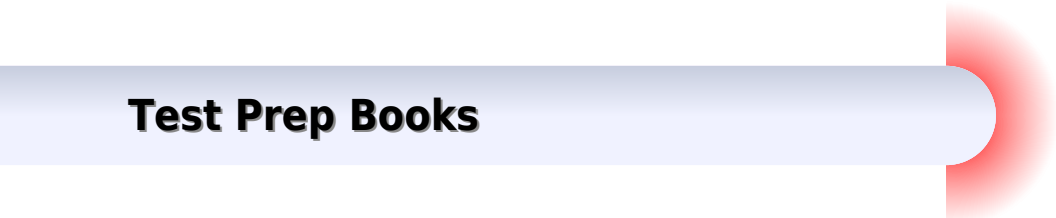
$$\Delta H^\circ = \sum n H^\circ(\text{products}) - \sum m H^\circ(\text{reactants})$$

$$\Delta S^\circ = \sum n S^\circ(\text{products}) - \sum m S^\circ(\text{reactants})$$

$$\Delta G^\circ = \sum n \Delta G^\circ_f(\text{products}) - \sum m \Delta G^\circ_f(\text{reactants})$$

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Table of Contents Practice Exam 3 Answer Key Emathinstructions

1. Understanding the eBook Practice Exam 3 Answer Key Emathinstructions
 - The Rise of Digital Reading Practice Exam 3 Answer Key Emathinstructions
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Exam 3 Answer Key Emathinstructions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Exam 3 Answer Key Emathinstructions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Exam 3 Answer Key Emathinstructions
 - Personalized Recommendations
 - Practice Exam 3 Answer Key Emathinstructions User Reviews and Ratings

- Practice Exam 3 Answer Key Emathinstructions and Bestseller Lists
- 5. Accessing Practice Exam 3 Answer Key Emathinstructions Free and Paid eBooks
 - Practice Exam 3 Answer Key Emathinstructions Public Domain eBooks
 - Practice Exam 3 Answer Key Emathinstructions eBook Subscription Services
 - Practice Exam 3 Answer Key Emathinstructions Budget-Friendly Options
- 6. Navigating Practice Exam 3 Answer Key Emathinstructions eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Exam 3 Answer Key Emathinstructions Compatibility with Devices
 - Practice Exam 3 Answer Key Emathinstructions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Exam 3 Answer Key Emathinstructions
 - Highlighting and Note-Taking Practice Exam 3 Answer Key Emathinstructions
 - Interactive Elements Practice Exam 3 Answer Key Emathinstructions
- 8. Staying Engaged with Practice Exam 3 Answer Key Emathinstructions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Exam 3 Answer Key Emathinstructions
- 9. Balancing eBooks and Physical Books Practice Exam 3 Answer Key Emathinstructions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Exam 3 Answer Key Emathinstructions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Exam 3 Answer Key Emathinstructions
 - Setting Reading Goals Practice Exam 3 Answer Key Emathinstructions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Exam 3 Answer Key Emathinstructions
 - Fact-Checking eBook Content of Practice Exam 3 Answer Key Emathinstructions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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