

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

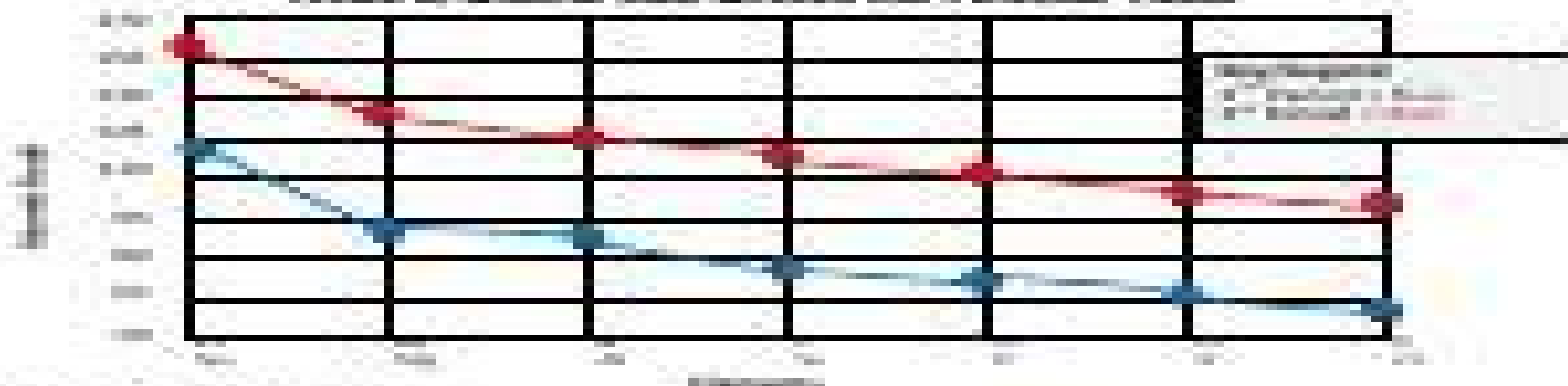
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of 5 data points (graph above)
2. Is a slight downward trend in this graph Period? If not, fill in the value graph
3. Labelled on box as "and trends" increased
4. decreased decreased

	1	2	3	4	5	6	7	8
1	1H	2He	3Li	4Be	5B	6C	7N	8O
2	9F	10Ne	11Na	12Mg	13Al	14Si	15P	16S
3	17Cl	18Ar	19K	20Ca	21Sc	22Ti	23V	24Cr
4	25Mn	26Fe	27Co	28Ni	29Cu	30Zn	31Ga	32Ge
5	33As	34Se	35Br	36Kr	37Rb	38Sr	39Y	40Zr
6	41Nb	42Mo	43Tc	44Ru	45Rh	46Pd	47Ag	48Cd
7	49In	50Sn	51Sb	52Te	53I	54Xe	55Cs	56Ba
8	57La	58Ce	59Pr	60Nd	61Pm	62Sm	63Eu	64Gd
9	65Tb	66Dy	67Ho	68Er	69Tm	70Yb	71Lu	72Hf
10	73Ta	74W	75Re	76Os	77Ir	78Pt	79Au	80Hg
11	81Tl	82Pb	83Bi	84Po	85At	86Rn	87Fr	88Ra
12	89Ac	90Th	91Pa	92U	93Np	94Pu	95Am	96Cm
13	97Bk	98Cf	99Es	100Fm	101Md	102No	103Lr	

Now label the trends in the graphing sheet

Trend in Atomic Size across the Periodic Table



Electronegativity Trend Equations

1. Following the trend in atomic size across the periodic table, decreased
2. This means that as you move from left to right, the electronegativity increases (the more electronegative the element, the more it attracts electrons)

Electronegativity Trend Equations

1. Electronegativity trend in the periodic table is increased (the more electronegative the element, the more it attracts electrons)
2. This means that as you move from left to right, the electronegativity increases (the more electronegative the element, the more it attracts electrons)

Periodicity Activity Okstate Answer Key

Huangqi Zhang



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