

Annotated periodic table of the elements

Group names*

- 1 Hydrogen & the alkali metals
- 2 Alkaline earth metals
- 3 Coinage metals
- 12 Volatile metals
- 13 Borogens
- 14 Chalcogens
- 15 Pnictogens
- 16 Chalcogens
- 17 Halogens
- 18 Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H 1.01 Hydrogen	2 He 4.003 Helium																
3 Li 6.94 Lithium	4 Be 9.012 Beryllium																
5 Na 22.99 Sodium	6 Mg 24.31 Magnesium																
7 K 39.10 Potassium	8 Ca 40.08 Calcium																
9 Rb 85.47 Rubidium	10 Sr 87.62 Strontium																
11 Cs 132.91 Cesium	12 Ba 137.33 Barium																
13 Fr 223.02 Francium	14 Ra 226.03 Radium																
		19 Sc 44.96 Scandium	20 Ti 47.87 Titanium	21 V 50.94 Vanadium	22 Cr 52.00 Chromium	23 Mn 54.94 Manganese	24 Fe 55.85 Iron	25 Co 58.93 Cobalt	26 Ni 58.69 Nickel	27 Cu 63.55 Copper	28 Zn 65.38 Zinc	29 Ga 69.72 Gallium	30 Ge 72.63 Germanium	31 As 74.92 Arsenic	32 Se 78.96 Selenium	33 Br 79.90 Bromine	34 Kr 83.80 Krypton
		35 Y 88.91 Yttrium	36 Zr 91.22 Zirconium	37 Nb 92.91 Niobium	38 Mo 95.94 Molybdenum	39 Tc 98.91 Technetium	40 Ru 101.07 Ruthenium	41 Rh 101.07 Rhodium	42 Pd 106.32 Palladium	43 Ag 107.87 Silver	44 Cd 112.41 Cadmium	45 In 114.82 Indium	46 Sn 118.71 Tin	47 Sb 121.76 Antimony	48 Te 127.60 Tellurium	49 I 126.91 Iodine	50 Xe 131.29 Xenon
		51 La 138.91 Lanthanum	52 Ce 140.12 Cerium	53 Pr 140.91 Praseodymium	54 Nd 144.24 Neodymium	55 Pm 144.91 Promethium	56 Sm 150.36 Samarium	57 Eu 151.96 Europium	58 Gd 157.25 Gadolinium	59 Tb 158.93 Terbium	60 Dy 162.50 Dysprosium	61 Ho 164.93 Holmium	62 Er 167.26 Erbium	63 Tm 168.93 Thulium	64 Yb 173.05 Ytterbium	65 Lu 174.97 Lutetium	66 Hf 178.49 Hafnium
		67 Ta 180.95 Tantalum	68 W 183.84 Tungsten	69 Re 186.21 Rhenium	70 Os 190.23 Osmium	71 Ir 192.22 Iridium	72 Pt 195.08 Platinum	73 Au 196.97 Gold	74 Hg 200.59 Mercury	75 Tl 204.38 Thallium	76 Pb 207.2 Lead	77 Bi 208.98 Bismuth	78 Po 209 Polonium	79 At 210 Astatine	80 Rn 222 Radon	81 Fr 223 Francium	82 Ra 226 Radium
		83 Ac 227.03 Actinium	84 Th 232.04 Thorium	85 Pa 231.04 Protactinium	86 U 238.03 Uranium	87 Np 237.05 Neptunium	88 Pu 244.06 Plutonium	89 Am 243.06 Americium	90 Cm 247.07 Curium	91 Bk 247.07 Berkelium	92 Cf 251.08 Californium	93 Es 252.08 Einsteinium	94 Fm 257.10 Fermium	95 Md 258.10 Mendelevium	96 No 259 Nobelium	97 Lr 260.10 Lawrencium	98 Rf 261.10 Rutherfordium
		99 Db 262.11 Dubnium	100 Sg 266.12 Seaborgium	101 Bh 264.10 Bohrium	102 Hs 277.10 Hassium	103 Mt 268.10 Meitnerium	104 Ds 271.10 Darmstadtium	105 Rg 272.10 Roentgenium	106 Cn 285.10 Copernicium	107 Nh 286.10 Nihonium	108 Fl 289.10 Flerovium	109 Mc 288.10 Moscovium	110 Lv 293.10 Livermorium	111 Ts 294.10 Tennessine	112 Og 294.10 Oganesson	113 Nh 294.10 Nihonium	114 Fl 294.10 Flerovium

Most common oxidation states, electronegativity (Allen-Pauling)

Atomic number and symbol, atomic weight and name

CLASSIFICATION

Metals	Nonmetals
Pre-transition	
Active	Corrosive
Alkali earth	
Alkaline	
Transition	Intermediate
Post-transition	
Metals	
Metals (Ru-Pt, Os-Au)	Metals gas

Active metals

Pre-transition metals

Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Fe, Mg and Al have any structural uses. Most form basic oxides (Be and B are amphoteric).

Rare earths

Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides

Soft, dense and reactive metals, those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form colored compounds.

Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals

Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalysts properties. Very weakly electropositive; inert to acids.

Frontier metals

Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and tellurates.

Metals

Nonmetals

Metalloids

Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals

B and Si are colorless gases, C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies; low to high electron affinities; moderate to high electronegativity values; relatively poor to moderately strong oxidising agents; tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (P₂O, CO, NO, N₂O).

*as black P, the most stable form in ambient conditions

Group 1

The starts as Li-Li, and then bifurcates into a La branch i.e. La-Ac, and a Lu branch i.e. Lu-Lu. Trends going down Lu-Lu, and the

• La branch, resembles those of groups 1–2

• Lu branch, more resembles those of groups 4–10

Corrosive nonmetals

O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid, I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases

Colorless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive

Approximation is likely to be either a metastable or a pure radioactive isotope

Periodic Table Unscrambled Table Answers

VM Jensen



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