

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	Across a period (→) atomic radius tends to _____. Across a period (→) first ionization energy tends to _____. Across a period (→) electronegativity tends to _____.						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	Down a Group (↓) atomic radius tends to _____. Down a Group (↓) first ionization energy tends to _____. Down a Group (↓) electronegativity tends to _____.						
Rb							
Cs							

Periodic Trends Worksheet With Graphs Answer Key

Wenbin Ji



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Table of Contents Periodic Trends Worksheet With Graphs Answer Key

1. Understanding the eBook Periodic Trends Worksheet With Graphs Answer Key
 - The Rise of Digital Reading Periodic Trends Worksheet With Graphs Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Worksheet With Graphs Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Worksheet With Graphs Answer Key
 - Personalized Recommendations
 - Periodic Trends Worksheet With Graphs Answer Key User Reviews and Ratings
 - Periodic Trends Worksheet With Graphs Answer Key and Bestseller Lists

5. Accessing Periodic Trends Worksheet With Graphs Answer Key Free and Paid eBooks
 - Periodic Trends Worksheet With Graphs Answer Key Public Domain eBooks
 - Periodic Trends Worksheet With Graphs Answer Key eBook Subscription Services
 - Periodic Trends Worksheet With Graphs Answer Key Budget-Friendly Options
6. Navigating Periodic Trends Worksheet With Graphs Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Worksheet With Graphs Answer Key Compatibility with Devices
 - Periodic Trends Worksheet With Graphs Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Worksheet With Graphs Answer Key
 - Highlighting and Note-Taking Periodic Trends Worksheet With Graphs Answer Key
 - Interactive Elements Periodic Trends Worksheet With Graphs Answer Key
8. Staying Engaged with Periodic Trends Worksheet With Graphs Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Worksheet With Graphs Answer Key
9. Balancing eBooks and Physical Books Periodic Trends Worksheet With Graphs Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Worksheet With Graphs Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Worksheet With Graphs Answer Key
 - Setting Reading Goals Periodic Trends Worksheet With Graphs Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Worksheet With Graphs Answer Key
 - Fact-Checking eBook Content of Periodic Trends Worksheet With Graphs Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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