

Quick Lab

Purpose To use a graph to identify period and group trends

Materials

- graph paper
- pencil



Periodic Trends in Ionic Radii

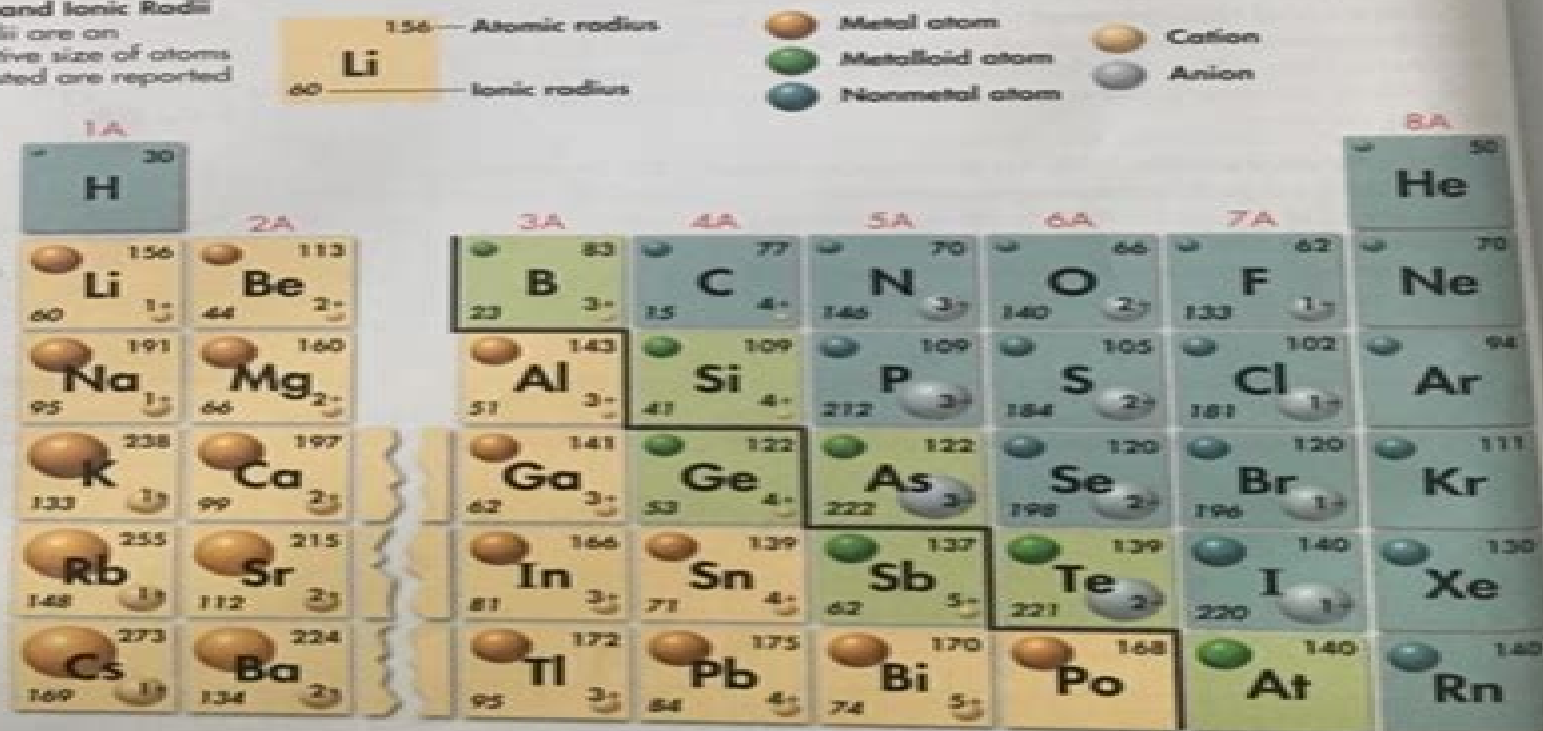
Procedure

Use the data presented in Figure 6.23 to plot ionic radius versus atomic number.

Analyze and Conclude

- 1. Compare** How does the size change when an atom forms a cation and when an atom forms an anion?
- 2. Describe** How do the ionic radii vary within a group of metals? How do they vary within a group of nonmetals?
- 3. Describe** What is the shape of a portion of the graph that corresponds to one period?
- 4. Compare and Contrast** Is the trend across a period similar or different for Periods 2, 3, 4, and 5?
- 5. Explain** Propose explanations for the trends you have described for ionic radii within groups and across periods.

Figure 6.23 Atomic and Ionic Radii
Atomic and ionic radii are an indication of the relative size of atoms and ions. The data listed are reported in picometers (pm).



Periodic Trends In Ionic Radii Lab Answers

N Colangelo



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Elements of the P Block Charlie Harding, David Arthur Johnson, Rob Janes, 2002 This book covers the chemistry of the non metallic elements the halogens boron carbon nitrogen oxygen silicon phosphorus and sulfur and uses their role in agriculture for example nitrogen and sulfur industry for example sulfuric acid and everyday life for example the chlorination of drinking water to illustrate this chemistry Their role in organic chemistry and biochemistry is also emphasized Two interactive CD ROMs accompany the book incorporating electronic questions that facilitate revision consolidation This book is part of The Molecular World series which aims to provide a broad foundation in chemistry **Cumulated Index Medicus**, 1964

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