

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Periodic Trends And The Properties Of Elements Lab Answers

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Periodic Trends And The Properties Of Elements Lab Answers:

Lab Experiments in Introductory Chemistry Phil Reedy, Donald J. Wink, Sharon Fetzner-Gislason, 2003-03-21 The manual contains laboratory experiments written specifically for the prep chem lab as well as for the general chemistry course Available as a complete manual or custom published at <http://custompub.whfreeman.com> **Contemporary Chemistry** Leonard Saland, 1993 This comprehensive guide gives you lesson plans activities and tests for two sequential semester long chemistry courses It is designed to work with our student book Contemporary Chemistry Each lesson plan features a DO NOW section to engage students as soon as they get to class instructional objectives an aim for that class period a motivational application questions or demonstrations to help students draw valid conclusions homework assignments You also get term calendars weekly tests and complete answer keys *Laboratory Manual for Principles of General Chemistry* Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures *Prentice Hall Science Explorer* Michael J. Padilla, 2002

Chemistry McGraw-Hill Staff, 2001-07 **Laboratory Manual for Principles of General Chemistry** J. A. Beran, Mark Lassiter, 2022-08-16 *Laboratory Manual for Principles of General Chemistry* 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section **Hands-On General Science Activities With Real-Life Applications** Pam Walker, Elaine Wood, 2008-04-21 In this second edition of *Hands On General Science Activities with Real Life Applications* Pam Walker and Elaine Wood have completely revised and updated their must have resource for science teachers of grades 5-12 The book

offers a dynamic collection of classroom ready lessons projects and lab activities that encourage students to integrate basic science concepts and skills into everyday life

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Chemistry John S. Phillips,Cheryl Wistrom,2000 Exploring General Chemistry in the Laboratory Colleen F. Craig,Kim N. Gunnerson,2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

Guided Inquiry Experiments for General Chemistry Nancy K. Kerner,Ram S. Lamba,2007-10-19 The use of the laboratory is a valuable tool in developing a deeper understanding of key chemical concepts from the experimental process This lab manual encourages scientific thinking enabling readers to conduct investigations in chemistry It shows how to think about the processes they are investigating rather than simply performing a laboratory experiment to the specifications set by the manual Each experiment begins with a problem scenario and ends with questions requiring feedback on the problem

Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 Laboratory Manual Jo A. Beran,1990 *Science Spectrum* Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2003-03

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Chemical Principles, Properties, and Reactions Kenneth L. Henold,Frank Walmsley,1984 *Laboratory Exercises for Preparatory Chemistry* Kathy Dodds Tyner,1994-06 Laboratory Exercises for Preparatory Chemistry is the perfect complement to a one semester preparatory chemistry laboratory course Tyner s manual emphasizes the application of chemistry and the principles of science to everyday life The labs are directly applicable to the

real world and often contain supplemental assignments that illustrate an application The Software Encyclopedia ,1988

Parallel Curriculum Units for Science, Grades 6-12 Jann H. Leppien, Jeanne H. Purcell, 2011-02-15 Teachers at various stages of professional development in curriculum design will find these materials powerful examples that will guide their growth and development and sharpen their skills Mary L Slade Professor James Madison University Harrisonburg VA Supercharge your science curriculum to challenge all students Based on the best selling book The Parallel Curriculum this professional development resource gives multifaceted examples of rigorous learning opportunities for science students in Grades 6 12 The four sample units revolve around genetics the convergence of science and society the integration of language arts and biology and the periodic table The editors and contributors provide user friendly methods for creating more thoughtful lessons and show how to differentiate them for the benefit of all students Included are field tested and standards based strategies that guide students through Exploring the nature of knowledge Discovering connections between science and other subjects Deepening science comprehension according to their interests and abilities Connecting science to society through the study of genetics historic events literature and chemistry Each unit includes subject matter background a content framework study components teacher reflections and sample lessons Also available are online content tools such as handouts PowerPoint presentations and research activities Breathe new life into science learning with this powerful guidebook written by master educators *The Science Teachers Bulletin* ,1989

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