

2.6 - Periodic Trends worksheet

Name _____

I. Short answer

1. Label each group on the periodic table.
2. Explain what valence electrons are.
3. Explain what makes an element stable?

II. Answer the questions using the symbols on the periodic table

1. What elements are in period 5?

What element(s) is a

2. halogens

3. lanthanides

alkaline earth metals

Snoble gases

Transition metals

Alkali metals

[illegible]

What elements are in the sublevel

98

107

Age Group	Percentage of Respondents
18-24	~15%
25-34	~25%
35-44	~10%
45-54	~20%
55-64	~30%
65+	~40%

126

What elements are

13. metals.

14. nonmetals

15. metalloids,

16. Which element is the most reactive nonmetal?

17. Which element is the most reactive metal?

18. Which element below is the least reactive?

19. Which element has the largest radius in period 4?

20. Which element has the largest radius in the Alkali metals family?

21. Which element has the highest ionization energy in period four?

22. Which elements has the second largest atomic number?

23. Which element has the largest atomic mass?

24. Which element is located at period 4, group 15?

25. Which element is located at period 5, group 2?

27 Which elements will form +1 ions?

28. Which group of metals will have the same properties

29 Which group of nonmetals will have the same properties

30. On the back, define all chapter 5, section 3 vocabulary terms.

[illegible][illegible]

Periodic Table Trends Activities Worksheet

George Graybill

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

Periodic Table Trends Activities Worksheet:

ChemDiscovery Teacher Edition Olga I. Agapova, 2002 **Trends in Teaching Experimentation in the Life Sciences** Nancy J. Pelaez, Stephanie M. Gardner, Trevor R. Anderson, 2022-05-11 This book is a guide for educators on how to develop and evaluate evidence based strategies for teaching biological experimentation to thereby improve existing and develop new curricula It unveils the flawed assumptions made at the classroom department and institutional level about what students are learning and what help they might need to develop competence in biological experimentation Specific case studies illustrate a comprehensive list of key scientific competencies that unpack what it means to be a competent experimental life scientist It includes explicit evidence based guidelines for educators regarding the teaching learning and assessment of biological research competencies The book also provides practical teacher guides and exemplars of assignments and assessments It contains a complete analysis of the variety of tools developed thus far to assess learning in this domain This book contributes to the growth of public understanding of biological issues including scientific literacy and the crucial importance of evidence based decision making around public policy It will be beneficial to life science instructors biology education researchers and science administrators who aim to improve teaching in life science departments Chapters 6 12 14 and 22 are available open access under a Creative Commons Attribution 4 0 International License via link [springer.com](https://www.springer.com) *Hands-On General Science Activities With Real-Life Applications* Pam Walker, Elaine Wood, 2008-04-21 In this second edition of Hands On General Science Activities with Real Life Applications Pam Walker and Elaine Wood have completely revised and updated their must have resource for science teachers of grades 5 12 The book offers a dynamic collection of classroom ready lessons projects and lab activities that encourage students to integrate basic science concepts and skills into everyday life Who's the New Kid in Chemistry? John D. Butler, 2013-12-12 Who s the New Kid in Chemistry offers an unprecedented look at student engagement and teacher best practices through the eyes of an educational researcher enrolled as a public high school student Over the course of seventy nine consecutive days John D Butler participates in and observes Rhode Island 2013 Teacher of the Year Jessica M Waters s high school chemistry class documenting his experiences as they unfold Who s the New Kid in Chemistry is a compelling example of what can be accomplished when an educational researcher and teacher collaborate in the classroom This work includes a discussion on flexible homework assignments data driven instruction and thirty teacher best practices This book is an invaluable resource for teachers across all content areas masters and doctoral research method classes and future Teachers of the Year **MnM_POW-Science-PM-10 (Updated)** Vibha Arora, Anju Sachdeva, Sushma Sardana, MnM_POW Science PM 10 Updated **The Science Teachers Bulletin** ,1989 **Census Bureau Education Program, Update 1** ,1992 Lists Bureau of the Census resources of interest to teachers in grades K 12 School Library Journal ,1988 The Software Encyclopedia ,1997 *Atoms, Molecules & Elements: The Periodic Table Gr. 5-8* George Graybill, 2015-10-01 This is the chapter slice The Periodic Table from the full lesson plan Atoms Molecules

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Periodic Table Cindy Blobaum,2005 What do chocolate chip cookies chemistry and logic have in common They are the basis for a unit that lets students become actively engaged in discovering the arrangement of the periodic table This learning activity takes the periodic table out of the static presentation usually associated with textbooks and chemistry courses and interjects an element of discovery The two activities in this unit provide students with information that they have to arrange in organized

charts In the process of creating the arrangements students will be involved in problem solving and will gain an appreciation for the scientific process of exploration and verification This dynamic unit meets national science standards in seven teaching and content areas Bring the periodic table to life with this hands on minds on unit Book jacket **Simply Science** ,1997

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attempts have not always been successful Eventually the form of the periodic table proposed by D I Mendeleev became the most common and accepted which is still used today It is known that chemical elements are located in the periodic system according to certain patterns taking into account the charge of the atomic nucleus atomic mass electronegativity ionization energy the metallic nature of the element and the basicity of oxides Understanding the trends in these characteristics in the periodic system allows you to predict the properties of a particular element and also provide for the course of certain chemical reactions The special feature of the book is that it describes the astrophysics of elements and also provides basic information about reactions that take place in stars or interstellar space The book is not overloaded with chemical equations so it is suitable for those who want to acquire fundamental knowledge about the elements as well as for those who want to expand their horizons

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