



Name: ANSWER KEY
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THE PERIODIC TABLE

Review Worksheet

1. Fill in the blanks with the terms and words most appropriate to complete each sentence.

- a) In the mid-nineteenth century, only sixty or so elements had been discovered. In 1869 (a year) a Russian chemist and inventor named Dmitri Mendeleev created a chart called the periodic table that predicted the properties of elements that had not yet been discovered. He did this by arranging the elements according to their increasing atomic numbers and by their physical and chemical properties.
- b) Each element in the periodic table is represented by an element box. In this box there is the following information: two numbers, one is a whole number which is called the atomic number, and the other number is larger and contains a decimal, and this one is called the atomic mass. In the box, there is also the short form of the element called the element symbol, and then the full identity of the element which is written using the element name.
- c) Today the modern periodic table is arranged into columns and rows. The columns are called groups and there are 18 (a number) of them. The rows are called periods and there are 7 (a number) of them. The periodic table is divided by a zigzag line also known as the staircase that separates the metals, found on the left of it, from the non-metals, found on the right of it. Directly around the zigzag line are eight elements called metalloids.
- d) All the metals, except for mercury, which is a liquid, are found as solids at room temperature. The metals are also mostly shiny in color, are ductile, malleable, conduct electricity and some are attracted to magnets.
- e) At room temperature, most of the non-metals are found in the gas state except for bromine which is a liquid, and iodine selenium carbon, phosphorus and sulfur which are all found as solids. These elements are colorless or come in various colors, are not ductile, not malleable and do not conduct electricity nor are they attracted to magnets.
- f) Metalloids are elements that have some properties of metals and some properties of non-metals. They are all found in the solid state. Most have a silvery metallic luster, they are brittle and are medium to good electrical conductors.
- g) Chemical families are groups on the periodic table that contain elements that share similar properties.
- h) Group 1 elements are called the alkali metals and they contain the most reactive group of all the metals. They react very violently with water and therefore must be stored under mineral oil so that the water vapor in the air cannot react with these elements. These metals get more reactive as you go down (up or down) the group and the most reactive of all the elements in this group is francium.
- i) Group 2 elements are called the alkaline earth metals and they contain the second most reactive group of all the metals. They also react with water and get more reactive as you go down (up or down) the group.

Periodic Table Review Sheet Answers

Leo J. Malone, Theodore O. Dolter



Periodic Table Review Sheet Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03
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extensive coverage of the principles but there is far less emphasis on the skills needed to actually solve problems

Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors' approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors' philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors' approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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web compiled by g izquierdo rodríguez page 2 1 directly proportional to the product of their charges $f \propto q_1 q_2$ 2 inversely proportional to square of the distance between them $f \propto \frac{1}{r^2}$ therefore $f \propto \frac{q_1 q_2}{r^2}$ coulomb s law in symbols $F = k \frac{q_1 q_2}{r^2}$ coulomb s law the magnitude of the electrostatic force exerted by one point charge q

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calculate the number of electrons that must be transferred to produce this charge

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web choose 1 answer the solution becomes negatively charged due to the majority Cl⁻ ions a the solution becomes negatively charged due to the majority Cl⁻ ions the solution becomes positively charged due to the stronger Mg²⁺ ions b the solution becomes positively charged due to the stronger Mg²⁺ ions the solution remains neutrally

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web answer d somewhere in region d justification with the test charge and c 1 being negative there is a repulsive force on the test charge to the right from c 2 there is an attractive force on the test charge to the left by referring to coulomb's law □□

□□□□ 1 □□ 2 □□ 2 we know that the force

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