

Name: **Brittany French**

Date: **October 4, 2012**

## Student Exploration: Electron Configuration

### Activity B: Atomic radii

#### Get the Gizmo ready:

- On the PERIODIC TABLE tab, select **Na** (sodium).
- Select the ELECTRON CONFIGURATION tab.



**Question:** How do the radii of atoms change across a period of the periodic table?

1. **Predict:** Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the **atomic radii** will change as electrons are added to a shell?

**It will decrease**

2. **Arrange:** Create a proper electron configuration for sodium. After clicking **Check**, note the **Electron configuration** and the **Atomic radius** now listed at right.

Sodium electron configuration: **1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>1</sup>**

Atomic radius: **190 Pico meters**

3. **Compare:** Click **Next element**, and then add an electron to the magnesium atom. Click **check**, and record the electron configuration and atomic radius below.

Magnesium electron configuration: **1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup>**

Atomic radius: **145 Pico meters**

4. **Gather data:** Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is *pm*.)

Element	Number of electrons	Electron configuration	Atomic radius (pm)
Aluminum	13	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>1</sup>	118
Silicon	14	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>2</sup>	111
Phosphorus	15	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>3</sup>	98
Sulfur	16	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>4</sup>	88
Chlorine	17	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>5</sup>	79
Argon	18	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> 3p <sup>6</sup>	71

# Periodic Table Exploration Answer Key

**ML Yell**



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