

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- $P_1 V_1 = P_2 V_2$

Charles' Theory and Boyle's Law

- Expansion of gases is caused by the activity of the gas molecules impinging the container.
- If the gas is compressible to small the volume it has large pressure exerts the pressure on the walls inside.
- The gas exerts pressure on walls of the container.
- Expansion of the gas will reduce.

Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. After the volume is reduced, it maintains a volume of 600 ml. above the pressure inside of 800 kPa. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad 800 \text{ kPa} \times 800 \text{ ml} = P_2 \times 600 \text{ ml}$$

$$P_2 = \frac{800 \text{ kPa} \times 800 \text{ ml}}{600 \text{ ml}} = 1066.67 \text{ kPa}$$

Ex. 2. A gas is pressure of 100 kPa and has a volume of 500 ml. If the pressure is increased until the volume is 400 ml, what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad 100 \text{ kPa} \times 500 \text{ ml} = P_2 \times 400 \text{ ml}$$

$$P_2 = \frac{100 \text{ kPa} \times 500 \text{ ml}}{400 \text{ ml}} = 125 \text{ kPa}$$

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