

Chemical Formula

Talking about proportions molecules and relative of atoms present (e.g. hydrogen)

Molecular $\text{C}_{10}\text{H}_{12}\text{O}_4$ (molecular)

Formula represents the number of atoms of each
in a single molecule

Empirical $\text{C}_5\text{H}_6\text{O}_2$ (molecular)

Formula represents the simplest ratio of
positive and negative ions in a formula

Atomic Mass

The average mass of atoms present in a sample

Mass

$\text{C}_{10}\text{H}_{12}\text{O}_4$ mass of one molecule

$\text{C}_5\text{H}_6\text{O}_2$ mass of one ion

Substance

$\text{C}_{10}\text{H}_{12}\text{O}_4$ substance

$\text{C}_5\text{H}_6\text{O}_2$ substance

$\text{C}_{10}\text{H}_{12}\text{O}_4$

Note Taking Guide Episode 601 Physics

Clifford Lane



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