

A. Choose the correct answer

- _____ 1. Equilibrium is characterized by
 - a. opposing processes occurring at equal rates.
 - b. an open system.
 - c. net change in the amount of substance in one phase.
 - d. changes in physical states.
- _____ 2. The process of changing from a solid to a gas is known as
 - a. condensation.
 - b. sublimation.
 - c. deposition.
 - d. vaporization.
- _____ 3. Which of the following liquids is most volatile?
 - a. water
 - b. ether
 - c. molten salt
 - d. skim milk
- _____ 4. A phase is defined as any part of a system that has
 - a. the same composition throughout.
 - b. the same properties throughout.
 - c. the same properties and composition throughout.
 - d. the same temperature throughout.
- _____ 5. Dry ice sublimates into carbon dioxide gas. If the proper conditions are maintained and the system is closed, the dry ice and the carbon dioxide gas will eventually
 - a. become the same phase.
 - b. reach equilibrium.
 - c. stop changing phase.
 - d. Both (a) and (b)
- _____ 6. If the vapor pressure of water is 4.25 kPa at 30.0°C, the vapor pressure of water at 10.0°C is most likely to be
 - a. 0 kPa.
 - b. 1.23 kPa.
 - c. 4.25 kPa.
 - d. 12.34 kPa.
- _____ 7. The temperature at which the vapor pressure of a substance equals atmospheric pressure is its
 - a. boiling point.
 - b. critical point.
 - c. melting point.
 - d. triple point.
- _____ 8. The amount of energy needed to melt one mole of a substance is its molar
 - a. enthalpy of fusion.
 - b. enthalpy of vaporization.
 - c. entropy of fusion.
 - d. entropy of vaporization.

Prentice Hall Chemistry Standardized Test Questions

Intermolecular Forces

H Kauffman



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