

CHEMISTRY SEMESTER 1 REVIEW QUESTIONS KEY

These questions are a small selection for each unit this semester. Refer to the exam scoring-guide and your self-reflection guide to determine which questions you should focus your studies on. Answers to select questions can be found at the end of this document.

Unit 1: Gas Laws

1. A gas is heated in a closed, rigid vessel. What happens to the molecules inside?
The molecules start to move faster. They hit the inside of the container with greater force, which increases the pressure.
2. A gas is compressed (volume decreased) at a constant temperature. What happens to the pressure inside the container? Explain your reasoning.
Since the volume is decreased, the molecules hit the sides of the container with increased frequency. This raises the pressure.
3. When inflating a car tire, the pressure doubles from 200 to 400 kPa. Why would the pressure double when the volume is increasing?
The volume may be increasing, but it is because more molecules are being added to the system. More molecules means more collisions, which increases the pressure.
4. What is the relation between temperature and pressure when volume is held constant?
With a constant volume, if temperature increases pressure increases as well. Think about heating an aerosol can (hair spray) – the volume is rigid and fixed. If you heat the can, the internal molecules move faster, hitting the sides harder and more frequently. This will increase the pressure so much that the can may explode.

Unit 2: Atomic structure and theory, periodic table

5. How many protons, neutron, and electrons are in a neutral atom of Ag-108?
Protons = 47
Neutrons = 61
Electrons = 47

6. Fill in the table below.

Name	Sodium	Nitrogen
Symbol	$^{23}_{11}\text{Na}$	$^{15}_7\text{N}$
Atomic number	11	7
Mass number	23	15
Protons	11	7
Neutrons	12	8
Electrons in a neutral atom	11	7

7. Analyze the following two isotopes. What is the same between both isotopes? What is different? $^{15}_7\text{N}$ and $^{14}_7\text{N}$
Both isotopes have the same number of protons (7) yet a different number of neutrons (8 and 7, respectively). The different number of neutrons yields a different mass number for the isotopes.
8. Chlorine has 2 isotopes, Cl-35 and Cl-37. Cl-35 has an abundance of 75.77% while Cl-37 has an abundance of 24.23%. Will the atomic mass of chlorine be closer to 35 or 37? Explain your reasoning?
Atomic mass will be closer to Cl-35 because 'atomic mass' refers to the average mass of all isotopes of an element. If most chlorine isotopes have a mass of 35 while only a few have a mass of 37, the average mass will be closer to 35.
9. Draw the Bohr model for sulfur. Label protons, neutrons, electrons, nucleus, and energy levels.



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Table of Contents Prentice Hall Chemistry Chapter 1review Answers

1. Understanding the eBook Prentice Hall Chemistry Chapter 1review Answers
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 1review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 1review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 1review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 1review Answers
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 1review Answers User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 1review Answers and Bestseller Lists
5. Accessing Prentice Hall Chemistry Chapter 1review Answers Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 1review Answers Public Domain eBooks
 - Prentice Hall Chemistry Chapter 1review Answers eBook Subscription Services
 - Prentice Hall Chemistry Chapter 1review Answers Budget-Friendly Options

6. Navigating Prentice Hall Chemistry Chapter 1review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 1review Answers Compatibility with Devices
 - Prentice Hall Chemistry Chapter 1review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 1review Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 1review Answers
 - Interactive Elements Prentice Hall Chemistry Chapter 1review Answers
8. Staying Engaged with Prentice Hall Chemistry Chapter 1review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 1review Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 1review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 1review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 1review Answers
 - Setting Reading Goals Prentice Hall Chemistry Chapter 1review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 1review Answers
 - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 1review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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