

13.1

THE NATURE OF GASES

P.1

Section Review

Objectives

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

Vocabulary

- | | | |
|------------------|------------------------|-----------------------------|
| • kinetic energy | • vacuum | • pascal (Pa) |
| • kinetic theory | • atmospheric pressure | • standard atmosphere (atm) |
| • gas pressure | • barometer | |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The kinetic theory describes the 1 of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly 2, that the particles of gas are relatively 3, move 4 of each other, and are in constant 5 motion. The 6 between particles are perfectly elastic so that the total 7 remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure 8 pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.

- motion
- empty space
- far apart
- independently
- random / rapid
- collisions
- Kinetic energy
- atmospheric
- 0°C / 273 K
- 1 atm / 101.3 kPa

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

ST 11. Atmospheric pressure is 760 mm Hg.

AT 12. The SI unit of pressure is the pascal.

↳ a multiple of kilo pascal

18 pts

- 2 pt if late
 - 1 pt if NO ID

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 - The Rise of Digital Reading Prentice Hall Chemistry Worksheet Answers
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 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Worksheet Answers
 - User-Friendly Interface
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 - Personalized Recommendations
 - Prentice Hall Chemistry Worksheet Answers User Reviews and Ratings
 - Prentice Hall Chemistry Worksheet Answers and Bestseller Lists

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 - Interactive Elements Prentice Hall Chemistry Worksheet Answers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Worksheet Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Worksheet Answers
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 - Creating a Diverse Reading Collection Prentice Hall Chemistry Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Worksheet Answers
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 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Worksheet Answers
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 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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