

6.9) It is often useful to write the equations for the reactions described:



The reaction for which the heat of reaction is requested is the formation of methane from its constituent elements.



The  $\Delta H$  for this reaction can be calculated by summing the  $\Delta H$ 's for a series of reactions which together add up to the same overall reaction (Hess's Law of Heat Summation).



These equations add up to the desired equation, so the  $\Delta H$  values are added:

$$\Delta H^\circ(\text{CH}_4) = +890. - 394 - 572 \text{ kJ/mol} = -76 \text{ kJ/mol}$$

6.10)  $\Delta H_{\text{fus}} = -924 \text{ kJ}$

6.11) The heat of reaction is the difference between the heat of formation of the products and the heat of formation of the reactants in the balanced equation:

$$\Delta H^\circ = 12 \times (-384) + 8 \times (-242) - 2 \times (+49) = -8278 \text{ kJ.}$$

There are two moles of benzene in the equation, so the heat released for one mole is 1/2 this amount, 3139 kJ/mol.

6.12)  $\Delta H_{\text{fus}} = -95.6 \text{ kJ}$

6.13) a. Heat lost =  $50. \text{ g} \times 0.129 \text{ J/g}^\circ\text{C} \times 68^\circ\text{C} = 260 \text{ J}$

b.  $q(\text{ice melting}) + q(\text{water warming}) = -q(\text{water cooling})$

$$5.0 \text{ g} \times 334 \text{ J/g} + 5.0 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 0^\circ\text{C}) = -100 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 70^\circ\text{C})$$

$$1670 \text{ J} + 20.92 T_f = -418.4 T_f + 29,288 \text{ J}$$

$$439.32 T_f = 27,618$$

$$T_f = 62.9 = 63^\circ\text{C}$$

6.14) a.  $q(\text{water}) = -q(\text{coins})$

$$c \times m \times \Delta T = -c \times m \times \Delta T$$

$$4.184 \text{ J/g}^\circ\text{C} \times 200. \text{ g} \times 5.9^\circ\text{C} = -c \times 113.4 \text{ g} \times (-74.1^\circ\text{C})$$

$$\text{Solving for } c \text{ gives: } 0.587 \text{ J/g}^\circ\text{C}$$

b. No, the specific heat capacity of pure gold is  $0.129 \text{ J/g}^\circ\text{C}$ .

6.15)  $\Delta H = 4.18 \text{ J/g}^\circ\text{C} \times 22.4 \text{ K} \times 70. \text{ g} = 6550 \text{ J}$

$$\text{For one mole of ammonium nitrate, } \Delta H = 6550 \times 60.05/20 = 26 \text{ kJ/mol.}$$

# Prentice Hall Chemistry Thermochemistry Answers

**Kathleen Armour**



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## **Table of Contents Prentice Hall Chemistry Thermochemistry Answers**

1. Understanding the eBook Prentice Hall Chemistry Thermochemistry Answers
  - The Rise of Digital Reading Prentice Hall Chemistry Thermochemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Thermochemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Prentice Hall Chemistry Thermochemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Thermochemistry Answers

- Personalized Recommendations
- Prentice Hall Chemistry Thermochemistry Answers User Reviews and Ratings
- Prentice Hall Chemistry Thermochemistry Answers and Bestseller Lists
- 5. Accessing Prentice Hall Chemistry Thermochemistry Answers Free and Paid eBooks
  - Prentice Hall Chemistry Thermochemistry Answers Public Domain eBooks
  - Prentice Hall Chemistry Thermochemistry Answers eBook Subscription Services
  - Prentice Hall Chemistry Thermochemistry Answers Budget-Friendly Options
- 6. Navigating Prentice Hall Chemistry Thermochemistry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Prentice Hall Chemistry Thermochemistry Answers Compatibility with Devices
  - Prentice Hall Chemistry Thermochemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Thermochemistry Answers
  - Highlighting and Note-Taking Prentice Hall Chemistry Thermochemistry Answers
  - Interactive Elements Prentice Hall Chemistry Thermochemistry Answers
- 8. Staying Engaged with Prentice Hall Chemistry Thermochemistry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Prentice Hall Chemistry Thermochemistry Answers
- 9. Balancing eBooks and Physical Books Prentice Hall Chemistry Thermochemistry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Prentice Hall Chemistry Thermochemistry Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Prentice Hall Chemistry Thermochemistry Answers
  - Setting Reading Goals Prentice Hall Chemistry Thermochemistry Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Prentice Hall Chemistry Thermochemistry Answers

- Fact-Checking eBook Content of Prentice Hall Chemistry Thermochemistry Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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