

Chemistry--Unit 10: Chemical Periodicity
Practice Problems

I. Classification of the elements

- 1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$ or abbreviated form*: $[\text{Ne}] 3s^2 3p^2$ *may use this form now!!!

- 2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$ or $[\text{Kr}] 5s^2 4d^{10} 5p^5$

- 3) Which family of elements is characterized by an $s^2 p^3$ configuration?

Nitrogen Family, or Group 15

- 4) Name the element that matches the following description:

- one that has 5 outer electrons on the third period of the periodic table **P, phosphorus**
- one with a $4s^2 4p^5$ electron configuration **Br, bromine**
- the 4th period Group 16 element **Se, selenium**

- 5) Identify the elements that have the following outer electron configurations?

- $2s^2 2p^4$ **O, oxygen**
- $4s^2$ **Ca, Calcium**
- $4s^2 3d^{10}$ **Zn, zinc**

- 6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

- 7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

- 8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

- 9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

- 10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

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EW Minium

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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Unveiling the Energy of Verbal Art: An Psychological Sojourn through **Prentice Hall Chemistry Practice Problems 6**

In a global inundated with displays and the cacophony of immediate connection, the profound power and mental resonance of verbal artistry frequently disappear into obscurity, eclipsed by the constant onslaught of noise and distractions. However, located within the lyrical pages of **Prentice Hall Chemistry Practice Problems 6**, a fascinating perform of literary splendor that impulses with organic emotions, lies an unique journey waiting to be embarked upon. Published by way of a virtuoso wordsmith, this enchanting opus courses visitors on a psychological odyssey, lightly revealing the latent potential and profound affect stuck within the delicate web of language. Within the heart-wrenching expanse of this evocative analysis, we shall embark upon an introspective exploration of the book is key styles, dissect their fascinating publishing type, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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