

Predicting Products of Chemical Reactions

This worksheet is designed to help you predict products of simple reactions of the four basic reaction types (synthesis, decomposition, single replacement, and double replacement) and combustion reactions.

For the first few reactions, the type of reaction is listed, you should predict the products, then balance. Further questions just have the reactants listed and you should decide on the type of reaction, as well as the correct products. Many of these reactions fall into the category of redox reactions, though do not let that confuse you...each can be described in terms of the four basic reaction types (except the combustion reactions).

Although states (s, l, g, aq) of the reactants and products are very important in a chemical reaction, don't worry about determining those for these problems. Rather, focus on what products might result from the reactants given. Pay particular attention to the ionic charge of species that you know forms ions with only one possible charge (e.g., alkali metals, alkaline earth metals, halogens, etc.)

1. Combustion: $\text{C}_2\text{H}_6 + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
2. Combustion: $2 \text{C}_4\text{H}_{10} + 11 \text{O}_2 \rightarrow 8 \text{CO}_2 + 6 \text{H}_2\text{O}$
3. Combustion: $\text{C}_5\text{H}_{12}\text{O}_2 + 7 \text{O}_2 \rightarrow 6 \text{CO}_2 + 5 \text{H}_2\text{O}$
4. Synthesis: $\text{Mg} + \text{I}_2 \rightarrow \text{MgI}_2$
5. Double displacement: $\text{CuCl}_2 + \text{H}_2\text{S} \rightarrow \text{CuS} + 2 \text{HCl}$
6. Double displacement: $\text{NaOH} + \text{HClO}_4 \rightarrow \text{NaClO}_4 + \text{H}_2\text{O}$
7. Decomposition: $\text{ZnCO}_3 + \text{heat} \rightarrow \text{ZnO} + \text{CO}_2$
8. Single replacement: $2 \text{HCl} + \text{Zn} \rightarrow \text{ZnCl}_2 + \text{H}_2$
9. single $2 \text{Na} + \text{MgCl}_2 \rightarrow 2 \text{NaCl} + \text{Mg}$
10. double $\text{CaCl}_2 + \text{K}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + 2 \text{KCl}$
11. synthesis $2 \text{K} + \text{Cl}_2 \rightarrow 2 \text{KCl}$
12. double $3 \text{BaCl}_2 + 2 \text{K}_3\text{PO}_4 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6 \text{KCl}$
13. double $\text{H}_2\text{SO}_4 + 2 \text{KOH} \rightarrow \text{K}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
14. synthesis $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
15. double $\text{Pb}(\text{NO}_3)_2 + 2 \text{KOH} \rightarrow \text{Pb}(\text{OH})_2 + 2 \text{KNO}_3$
16. double $\text{H}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow 2 \text{HCl} + \text{BaSO}_4$
17. single $\text{Ca} + 2 \text{AgCl} \rightarrow \text{CaCl}_2 + 2 \text{Ag}$
18. double $\text{H}_3\text{PO}_4 + \text{FeBr}_3 \rightarrow 3 \text{HBr} + \text{FePO}_4$
19. synthesis $6 \text{Li} + \text{N}_2 \rightarrow 2 \text{Li}_3\text{N}$
20. double $2 \text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2 + 2 \text{H}_2\text{O}$

Predicting Products Of Chemical Equations Answers

Tillery

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

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Table of Contents Predicting Products Of Chemical Equations Answers

1. Understanding the eBook Predicting Products Of Chemical Equations Answers
 - The Rise of Digital Reading Predicting Products Of Chemical Equations Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Predicting Products Of Chemical Equations Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Predicting Products Of Chemical Equations Answers
 - Personalized Recommendations
 - Predicting Products Of Chemical Equations Answers User Reviews and Ratings
 - Predicting Products Of Chemical Equations Answers and Bestseller Lists

5. Accessing Predicting Products Of Chemical Equations Answers Free and Paid eBooks
 - Predicting Products Of Chemical Equations Answers Public Domain eBooks
 - Predicting Products Of Chemical Equations Answers eBook Subscription Services
 - Predicting Products Of Chemical Equations Answers Budget-Friendly Options
6. Navigating Predicting Products Of Chemical Equations Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Predicting Products Of Chemical Equations Answers Compatibility with Devices
 - Predicting Products Of Chemical Equations Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Predicting Products Of Chemical Equations Answers
 - Highlighting and Note-Taking Predicting Products Of Chemical Equations Answers
 - Interactive Elements Predicting Products Of Chemical Equations Answers
8. Staying Engaged with Predicting Products Of Chemical Equations Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Predicting Products Of Chemical Equations Answers
9. Balancing eBooks and Physical Books Predicting Products Of Chemical Equations Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Predicting Products Of Chemical Equations Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Predicting Products Of Chemical Equations Answers
 - Setting Reading Goals Predicting Products Of Chemical Equations Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Predicting Products Of Chemical Equations Answers
 - Fact-Checking eBook Content of Predicting Products Of Chemical Equations Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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