Predicting Products of Chemical Reactions

This worksheet is designed to help you predict products of simple reactions of the four basic reaction types (synthesis, decomposition, single replacement, and double replacement) and combustion reactions.

For the first few reactions, the type of reaction is listed, you should predict the products, then balance. Further questions just have the reactions is listed and you should decide on the type of reaction, as well as the correct products. Many of these reactions fall into the category of redox reactions, though do not let that confuse you...each can be described in terms of the four basic reaction types (except the combustion reactions).

Although states (s, l, g, aq) of the reactants and products are very important in a chemical reaction, don't werry about determining those for these problems. Rather, focus on what products might result from the reactants given. Pay particular attention to the ionic charge of species that you know form ions with only one possible charge (e.g., alkali metals, alkaline earth metals, halogens, etc.)

20	Combustion:	C_nH_{12}	+	9 02		6002 + 6H20
2	Combustion:	2.C ₄ H ₄	+	14 O2		8 CO2 + 6 H20
3.	Combustion:	$C_0H_{10}O_7$		+ 7	Oz	→ 600, + 5H,0
4.	Synthesis:	Mg	+	I ₂	-	_MgI2
35.	Double displacement:	CuCl ₂	*	H ₃ S	\rightarrow	_ CuS + 2HC1
- 6.	Double displacement.	NaOH		HCIO ₄		NaClOg + HEO
7.	Decomposition:	ZnCO ₃		heat	-	Zn0 + Co.
(8)	Single replacement:	2.HCl		Zn	-•	ZnCl. + Hz
9.	sing le	2Na	+	$MgCl_2$	\rightarrow	2 Nacl + Mg
10.	double	Ca/Cl ₂	*	K ₂ CO ₂		Caco, - 2KCI
11.	synthesis	ZK	+	CI_2	→	2KC1
12.	double	3 BaCl,	+	2 K ₃ PO ₄	-+	Bas (PON) + GKC1
13.	double	H ₂ SO ₄	+	2 KOH	-	K2504 + 2H20
14.	Synthesis	4 AI	+	3 02		2 A1203
(15)	_domble	Pb(INO ₃)	la.	+ 2	кон	- Pb(OH) + 2KNO3
16.	double	H ₂ SO ₄	+	BaCl ₂	-+	2 HC1 + B & SO4
(17)	single	Ca	+	2 AgCI	-	Cacle + 2Ag
18.	double	H_bPO_4		$FeB\tau_{0}$	-	3 HBr + Fe PD4
(19.)	Synthesis	6 Li		MrNa		2 L: 3 N
20.	double	2 HCI		Mg(OH)	a.	→ MgCl2 + 2H20

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