

Precipitate Reactions: Chemical “Drop-Outs”

Many chemical reactions take place in solution. In this experiment, you will see some dissolved chemicals react when solutions are combined and drop out of solution or precipitate. In a chemical reaction, substances that react to form new substances are called reactants; new substances that are produced are called products. Substances that “fall out” of solution as solids, denoted $_{(s)}$, are called precipitates (ppt).

Goals:

- To learn about double replacement reactions.
- To learn solubility rules.
- To learn small-scale lab techniques.
- To write balanced chemical equations to describe reactions performed in lab.

Materials Needed:

- 6 small scale pipettes (empty, barium chloride, magnesium nitrate, copper (II) sulfate, potassium hydroxide, and sodium carbonate)
- reaction surface (Make sure the surface is clean, but DO NOT USE SOAP.)

Procedure:

1. Obtain materials.
2. Put 3 drops of sodium carbonate on the reaction surface in A. Add 3 drops of barium chloride solution. **Be careful to not cross contaminate the tip of the pipets.**
3. Record your observations in the data chart. Pay special attention to the color of each precipitate that forms, as well as its consistency (grainy, clumpy, etc.) . Use a Q-tip if needed to assist with your observations. Write “No Rxn” when no precipitate forms.
4. Make all of the remaining combinations listed in the data chart on the reaction surface in boxes B-J. Use 3 drops of each solution as you did in step 2. Record your observations in the data chart as soon as you make a combination.
5. Return all pipets **bulb down to the proper container** and clean the reaction surface with a paper towel.

Post-lab:

6. Complete the word equations for the reaction under part 2 in the equation chart.
7. Write the chemical equation for part 1 in the equation chart for each combination.
For your equations from part 1 in the equation chart:
8. If a precipitate formed, use the solubility chart to determine which of the products formed is the precipitate (use the solubility table). Circle the precipitate in the equation and label with $_{(s)}$. Label the other product with $_{(aq)}$.
9. If no precipitate is formed, label each product with $_{(aq)}$. Then, write “No Rxn” under the number in the pair column.
10. Answer post-lab questions.

Precipitate Lab Answers

John Kenkel



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