

**SUBJECT: SCIENCE**  
**CLASS : X**

**MAX. MARKS: 80**  
**DURATION: 3 HRS**

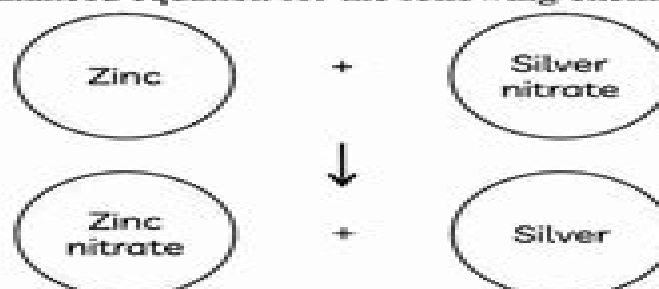
**General Instruction:**

1. This Question Paper has 5 Sections A-E.
2. **Section A** has 20 MCQs carrying 1 mark each.
3. **Section B** has 5 questions carrying 02 marks each.
4. **Section C** has 6 questions carrying 03 marks each.
5. **Section D** has 4 questions carrying 05 marks each.
6. **Section E** has 3 case based integrated units of assessment (04 marks each) with sub-parts of the values of 1, 1 and 2 marks each respectively.
7. All Questions are compulsory. However, an internal choice in 2 Qs of 5 marks, 2 Qs of 3 marks and 2 Questions of 2 marks has been provided. An internal choice has been provided in the 2marks questions of Section E
8. Draw neat figures wherever required. Take  $\pi = 22/7$  wherever required if not stated.

**SECTION – A**

**Questions 1 to 20 carry 1 mark each.**

1. Which of the following a balanced equation for the following chemical reaction:



- (a)  $\text{Zn} + \text{Ag}(\text{NO}_3)_2 \rightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$
  - (b)  $2\text{Zn} + 2\text{AgNO}_3 \rightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$
  - (c)  $\text{Zn} + 2\text{AgNO}_3 \rightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$
  - (d)  $\text{Zn} + \text{AgNO}_3 \rightarrow \text{Zn}(\text{NO}_3)_2 + \text{Ag}$
2. Which of the following will turn phenolphthalein pink?  
 (a)  $\text{NaOH}(\text{aq})$  (b)  $\text{HCl}(\text{aq})$  (c)  $\text{CH}_3\text{COOH}(\text{aq})$  (d)  $\text{H}_2\text{O}$
  3. While studying the saponification reaction, what do you observe when you mix an equal amount of colourless vegetable oil and 20% aqueous solution of  $\text{NaOH}$  in a beaker?  
 (a) The colour of the mixture has become dark brown.  
 (b) A brisk effervescence is taking place in the beaker.  
 (c) The outer surface of the beaker has become hot.  
 (d) The outer surface of the beaker has become cold.
  4. Why do we store silver chloride in dark-coloured bottles?  
 (a) To prevent precipitation of silver chloride  
 (b) To prevent decomposition of silver chloride  
 (c) To promote decomposition of silver chloride  
 (d) All of these

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