

FLAME TEST LAB PROCEDURE

Materials & Equipment

Safety Goggles

Gloves

Beakers, 250-mL, 2

Matches

Weighing dishes, 6

Wooden splints soaked in water, 6

Water, distilled or deionized, 250 mL

Calcium chloride, $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$, 1–1.5 g

Copper chloride, $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$, 1–1.5 g

Lithium chloride, LiCl , 1–1.5 g

Potassium chloride, KCl , 1–1.5 g

Sodium chloride, NaCl , 1–1.5 g

Strontium chloride, $\text{SrCl}_2 \cdot 6\text{H}_2\text{O}$, 1–1.5 g

Procedure:

1. Put on Safety Goggles.
2. Wear gloves when handling chemicals.
3. Obtain a 250-mL beaker about half-full with distilled or deionized water. Obtain six wooden splints that have been soaked in distilled or deionized water. Place them in this beaker of water to continue soaking at your lab station.
4. Fill a second 250-mL beaker about half-full with tap water. Label this beaker "rinse water".
5. Obtain six weighing dishes labeled " CaCl_2 ", " CuCl_2 ", " LiCl ", " NaCl ", " KCl ", " SrCl_2 ". Each contains a small scoopful (about 1–1.5 g) of each metallic solid in the corresponding weighing dish.
6. Dip the soaked end of one of the wooden splints in one of the metallic salts.
7. Light a match and then place the "dipped" wooden splint into the flame. Observe the color of the flame. Allow the splint to burn until the color fades. Try not to allow any of the solid to fall onto your desk. If necessary, repeat the test with the same splint and salt.
8. Immerse the wooden splint in the "rinse water" to fully extinguish it, as well as the match. Then discard both into the trash.
9. Record your observations for the flame color produced by the metallic salt in the Data Table.
10. Repeat Steps 6–8 for the other five metallic salts. Record your observations for the flame color produced by each metallic salt in the Data Table.
11. Clean up your area, put everything away, and wash your hands before leaving the lab.

Representative Wavelength, nm	Wavelength Region, nm	Color
410	400–425	Violet
470	425–480	Blue
490	480–500	Blue-green
520	500–560	Green
565	560–580	Yellow-green
580	580–585	Yellow
600	585–650	Orange
650	650–700	Red

Figure 1.

Prentice Hall Flame Test Lab 10 Answers

Raffaella Di Napoli



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Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

Prentice Hall Earth Science, 1988

Prentice Hall General Science: A voyage of discovery, 1992

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Applied Mechanics Reviews, 1964

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Mirror of Science and Art, 1915 Gas Journal, 1926 American Gas-light Journal and Chemical Repertory, 1910

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