

**Assessment #3 Review**

Name: \_\_\_\_\_

You must show all work for credit.

Due Date: \_\_\_\_\_ (No late packets will be accepted)

**I. Balance Equations***Balance the equations*

- $\text{Mg} + 2 \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + 2 \text{NaBr} \rightarrow \text{MgBr}_2 + 2 \text{Na}$
- $\text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
- $4 \text{P} + 5 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + 6 \text{KOH} \rightarrow 3 \text{K}_2\text{SO}_4 + 2 \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow \text{Fe} + 3 \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of  $\text{Fe}_2\text{O}_3$  in reaction #6? 3

*Write the balanced equation for problems # 8 and 9:*

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.  
 $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$

2 N on left, 2 N on right, 6 H on left, 6 H on right, so no N gets created or lost

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

moles

12. Using the balanced equation shown in problem #10, list the products of the reaction.

$\text{NH}_3$

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

$\text{N}_2$ ,  $\text{H}_2$

**II. Types of Reactions***Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



# Prentice Hall Chemistry Final Exam Review Answers

**L Reisser**



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