

Chemistry 105

Chapter 3 Review Problems

1. How many molecules are present in 180 grams of water? (The atomic weights of H and O are 1 g/mol and 16 g/mol, respectively.)
a) 6.0×10^{24}
b) 10
c) 5.6×10^{-23}
d) 1.8×10^7
e) 180×10^{23}
2. The atomic mass of sulfur is given as 32.07 g/mol. What is the mass of an average atom of sulfur?
a) 5.33×10^{-23} g
b) 32.07 g
c) 32.07×10^{-23} kg
d) 193×10^{-22} g
e) 6.02×10^{23} g
3. What is the mass of C present in 60 g of C_3H_4 ?
a) 12 g
b) 24 g
c) 48 g
d) 60 g
e) 6 g
4. What is the molar mass of C_3H_4O ?
a) 24 g/mol
b) 32 g/mol
c) 7 g/mol
d) 35 g/mol
e) 46 g/mol
5. What is the correct percentage by mass of nitrogen in HCN?
a) 33 %
b) 14%
c) 27%
d) 52%
e) 98%
6. What is the chief danger that is anticipated if CO_2 levels get too high in the atmosphere?
a) people will develop skin cancer from exposure to UV light
b) the atmosphere will develop an unpleasant odor
c) the temperature of the planet may climb by several degrees
d) plant life will be exterminated
e) The CO_2 will remove ozone from the stratosphere

Prentice Hall Chemistry Chapter 3 Practice Problems

Thomas Griffiths



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Table of Contents Prentice Hall Chemistry Chapter 3 Practice Problems

1. Understanding the eBook Prentice Hall Chemistry Chapter 3 Practice Problems
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 3 Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 3 Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 3 Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 3 Practice Problems
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 3 Practice Problems User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 3 Practice Problems and Bestseller Lists

5. Accessing Prentice Hall Chemistry Chapter 3 Practice Problems Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 3 Practice Problems Public Domain eBooks
 - Prentice Hall Chemistry Chapter 3 Practice Problems eBook Subscription Services
 - Prentice Hall Chemistry Chapter 3 Practice Problems Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Chapter 3 Practice Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 3 Practice Problems Compatibility with Devices
 - Prentice Hall Chemistry Chapter 3 Practice Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 3 Practice Problems
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 3 Practice Problems
 - Interactive Elements Prentice Hall Chemistry Chapter 3 Practice Problems
8. Staying Engaged with Prentice Hall Chemistry Chapter 3 Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 3 Practice Problems
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 3 Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 3 Practice Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 3 Practice Problems
 - Setting Reading Goals Prentice Hall Chemistry Chapter 3 Practice Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 3 Practice Problems
 - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 3 Practice Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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