

Chapter 7 Chemical Reactions

Section 7.2 Types of Reactions

(pages 199–205)

This section discusses how chemical reactions are classified into different types.

Reading Strategy (page 199)

Previewing. Skim the section and begin a concept map like the one below that identifies types of reactions with a general form.

As you read, add the general forms of each type of reaction. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



Classifying Reactions (pages 199–204)

- Name five general types of chemical reactions. Synthesis, decomposition, single replacement, double replacement, and combustion.
- Circle the letter of each equation that represents a synthesis reaction.
 - a. $2\text{Na} + \text{Cl}_2 \longrightarrow 2\text{NaCl}$
 - $2\text{NaCl} \longrightarrow 2\text{Na} + \text{Cl}_2$
 - $2\text{H}_2\text{O} \longrightarrow 2\text{H}_2 + \text{O}_2$
 - d. $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$
- Is the following sentence true or false? A decomposition reaction is the opposite of a synthesis reaction. true
- Write the equation for the decomposition of calcium carbonate into calcium oxide and carbon dioxide. $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$
- Circle the letter of the correct answer. Copper reacts with silver nitrate in a single-replacement reaction. What are the products of this reaction?
 - copper(II) nitride and silver oxide
 - b. copper(II) nitrate and silver
 - copper(II) oxide and silver nitrate
 - copper, nitrogen, and silver oxide

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