

## Study Guide

## Key Concepts



## 14.1 Properties of Gases

- Gases are easily compressed because of the space between particles in a gas.
- The amount of gas ( $n$ ), volume ( $V$ ), and temperature ( $T$ ) are factors that affect gas pressure ( $P$ ).



## 14.2 The Gas Laws

- As the pressure of a gas increases, the volume decreases, if the temperature is constant.
- As the temperature of an enclosed gas increases, the volume increases, if the pressure is constant.
- As the temperature of an enclosed gas increases, the pressure increases, if the volume is constant.

- The combined gas law allows you to do calculations for situations in which only the amount of gas is constant.



## 14.3 Ideal Gases

- To calculate the number of moles of a contained gas requires an expression that contains the variable  $n$ .
- Real gases differ most from an ideal gas at low temperatures and high pressures.



## 14.4 Gases: Mixtures and Movements

- In a mixture of gases, the total pressure is the sum of the partial pressures of the gases.
- Gases of lower molar mass diffuse and effuse faster than gases of higher molar mass.

## Vocabulary

- Boyle's law (p. 418)
- Charles's law (p. 420)
- combined gas law (p. 424)
- compressibility (p. 413)
- Dalton's law of partial pressures (p. 432)
- diffusion (p. 435)
- effusion (p. 435)
- Gay-Lussac's law (p. 422)
- Graham's law of effusion (p. 435)
- ideal gas constant (p. 426)
- ideal gas law (p. 426)
- partial pressure (p. 432)

## Key Equations

- Boyle's law:  $P_1 \times V_1 = P_2 \times V_2$
- Charles's law:  $\frac{V_1}{T_1} = \frac{V_2}{T_2}$
- Gay-Lussac's law:  $\frac{P_1}{T_1} = \frac{P_2}{T_2}$
- Combined gas law:  $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$
- Ideal gas law:  $P \times V = n \times R \times T$  or  $PV = nRT$
- Dalton's law:  $P_{\text{total}} = P_1 + P_2 + P_3 + \dots$
- Graham's law:  $\frac{\text{Rate}_A}{\text{Rate}_B} = \sqrt{\frac{\text{molar mass}_B}{\text{molar mass}_A}}$

## Organizing Information

Use these terms to construct a concept map that organizes the major ideas of this chapter.



**Concept Map 14** Solve the Concept Map with the help of an interactive guided tutorial.

with **ChemASAP**

Boyle's law

Charles's law

partial pressure

combined gas law

diffusion

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**AW Rasmussen**



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## **Table of Contents Prentice Hall Chemistry Chapter 14 Packet Answers**

1. Understanding the eBook Prentice Hall Chemistry Chapter 14 Packet Answers
  - The Rise of Digital Reading Prentice Hall Chemistry Chapter 14 Packet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 14 Packet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Prentice Hall Chemistry Chapter 14 Packet Answers
  - User-Friendly Interface

4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 14 Packet Answers
  - Personalized Recommendations
  - Prentice Hall Chemistry Chapter 14 Packet Answers User Reviews and Ratings
  - Prentice Hall Chemistry Chapter 14 Packet Answers and Bestseller Lists
5. Accessing Prentice Hall Chemistry Chapter 14 Packet Answers Free and Paid eBooks
  - Prentice Hall Chemistry Chapter 14 Packet Answers Public Domain eBooks
  - Prentice Hall Chemistry Chapter 14 Packet Answers eBook Subscription Services
  - Prentice Hall Chemistry Chapter 14 Packet Answers Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Chapter 14 Packet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Prentice Hall Chemistry Chapter 14 Packet Answers Compatibility with Devices
  - Prentice Hall Chemistry Chapter 14 Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 14 Packet Answers
  - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 14 Packet Answers
  - Interactive Elements Prentice Hall Chemistry Chapter 14 Packet Answers
8. Staying Engaged with Prentice Hall Chemistry Chapter 14 Packet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Prentice Hall Chemistry Chapter 14 Packet Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 14 Packet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 14 Packet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 14 Packet Answers
  - Setting Reading Goals Prentice Hall Chemistry Chapter 14 Packet Answers
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 14 Packet Answers
  - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 14 Packet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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