

CHEMICAL REACTIONS

Practice Problems

In your notebook, solve the following problems. Use the 3-step problem-solving approach you learned in Chapter I.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS

- Write the skeleton equation for the reaction between hydrogen and oxygen that produces water.
- Write the skeleton equation for the reaction that produces iron(II) sulfide from iron and sulfur.
- Write the skeleton equation representing the heating of magnesium carbonate to produce solid magnesium oxide and carbon dioxide gas.
- 4. Write a balanced equation for the production of HCl gas from its elements.
- Write a sentence that completely describes the chemical reaction represented by this balanced equation.

 $2HCl(aq) + CaCO_3(s) \rightarrow CO_2(g) + CaCl_2(aq) + H_2O(l)$

Write the word equation for the following equation. Write a sentence fully describing the reaction. Is the equation correctly balanced? Explain.

 $2Ag(s) + S(s) \rightarrow Ag_0S(s)$

- Write a balanced equation representing the formation of aqueous sulfuric acid from water and sulfur trioxide gas.
- 8. Write a balanced equation from this word equation.

aqueous silver nitrate + copper metal -- silver metal + aqueous copper nitrate

9. Write a balanced equation for the following word equation.

phosphorus + oxygen → tetraphosphorous decoxide

SECTION 11.2 TYPES OF CHEMICAL REACTIONS

- Write a balanced equation representing the reaction of magnesium with oxygen gas to produce magnesium oxide.
- Write the balanced equation for the reaction that occurs between aluminum and fluorine.
- Write the balanced equation for the production of oxygen gas and potassium chloride from the decomposition of potassium chlorate.
- Write the balanced equation for the reaction between hydrochloric acid and calcium metal. The products are hydrogen gas and calcium chloride.
- Write the balanced equation for the combustion of propane (C₃H₈) to produce carbon dioxide and water vapor.
- Write the balanced equation for the reaction between iron(III) chloride and sodium hydroxide. The products are iron(III) hydroxide and sodium chloride.

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