

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $\text{KClO}_3(\text{s})$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{23} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



Prentice Hall Chemistry Workbook Answers Chapter 12

Stoichiometry

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Table of Contents Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry

1. Understanding the eBook Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - The Rise of Digital Reading Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Personalized Recommendations
 - Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry User Reviews and Ratings
 - Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry and Bestseller Lists

5. Accessing Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry Free and Paid eBooks
 - Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry Public Domain eBooks
 - Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry eBook Subscription Services
 - Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry Compatibility with Devices
 - Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Highlighting and Note-Taking Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Interactive Elements Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
8. Staying Engaged with Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Setting Reading Goals Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Fact-Checking eBook Content of Prentice Hall Chemistry Workbook Answers Chapter 12 Stoichiometry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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