

Chemistry Chapter 8 Test

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- ____ 1. What type of bond would you expect in a compound of oxygen and sodium?
☒ a. ionic
 b. metallic
 c. covalent
 d. polar covalent
- ____ 2. The symbols δ^- and δ^+ indicate which of the following?
 a. electronegativity
 b. electrostatic force
 c. charge
☒ d. partial charge
- ____ 3. Which of the following represents a nonpolar covalent molecule?
 a. NaCl
☒ b. Br₂
 c. HCN
 d. Au
- ____ 4. The measure of an atom's ability to attract electrons to itself in a chemical bond is called _____.
 a. electron affinity
 b. polarity
☒ c. electronegativity
 d. ionization energy
- ____ 5. The polarity of a bond depends on the difference in the atoms' _____.
 a. electron affinities
☒ b. electronegativities
 c. atomic radii
 d. ionization energies
- ____ 6. What type of bond would you expect to form between nitrogen and potassium?
☒ a. ionic
 b. covalent
 c. polar covalent
 d. metallic
- ____  7. High electronegativity differences are found in _____ bonds.
 a. metallic
☒ b. ionic metal - nonmetal
 c. covalent
 d. Pauling
- ____  8. To achieve an octet when combining with potassium, would an atom of a Group 16 element be expected to lose, gain, or share electrons?
 a. lose
☒ b. gain
 c. share
☒ d. This cannot be determined from the given information. x
- ____ 9. Which of the following does not exist as a diatomic molecule?
☒ a. argon
 b. hydrogen
 c. chlorine
 d. oxygen
 H N O F Cl Br I
 Gas : H₂ N₂ O₂ F₂ Cl₂ Br₂ I₂
- ____ 10. Which of these is not a covalent molecule?
 a. F₂ =
 b. CO =
☒ c. LiF
 d. NH₃
- ____ 11. A dash in a Lewis structure represents _____.
☒ a. one electron
 b. two electrons
 c. four electrons
 d. unpaired electrons
- ____ 12. The central atom in a Lewis structure is usually the atom _____.

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