

## 11-3 Practice Problems

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1. Identify the limiting reactant when 1.22-g of  $\text{O}_2$  reacts with 1.65-g of  $\text{H}_2$  to produce water.

2. Identify the limiting reactant when 4.68-g of Fe reacts with 2.88-g of S to produce  $\text{FeS}$ .

3. Identify the limiting reactant when 2.87-g of  $\text{Mg}(\text{OH})_2$  reacts with 12.84-g of  $\text{HCl}$  to form  $\text{MgCl}_2$  and water.

4. Identify the limiting reactant when 6.25-g of  $\text{AgNO}_3$  reacts with 4.12-g of  $\text{NaCl}$  to form  $\text{NaNO}_3$  and  $\text{AgCl}$ .

5. Identify the limiting reactant when 7.81-g of  $\text{HCl}$  reacts with 5.24-g of  $\text{NaOH}$  to produce  $\text{NaCl}$  and  $\text{H}_2\text{O}$ .

6. Identify the limiting reactant when 6.33-g of  $\text{H}_2\text{SO}_4$  reacts with 5.92-g of  $\text{NaOH}$  to produce  $\text{Na}_2\text{SO}_4$  and water.

7. Identify the limiting reactant when 49.25-g of  $\text{CaC}_2$  reacts with 33.71-g of water to produce  $\text{Ca}(\text{OH})_2$  and  $\text{C}_2\text{H}_2$ .

8. Identify the limiting reactant when 63.14-g of  $\text{CaCl}_2$  reacts with 74.68-g of  $\text{Na}_2\text{CO}_3$  to produce  $\text{CaCO}_3$  and  $\text{NaCl}$ .

9. Identify the limiting reactant when 4.687-g of  $\text{SF}_4$  reacts with 6.281-g of  $\text{I}_2\text{O}_5$  to produce  $\text{IF}_3$  and  $\text{SO}_2$ .

10. If 4.1-g of Cr is heated with 9.3-g of  $\text{Cl}_2$ , what mass  $\text{CrCl}_3$  will be produced?

11. What mass of  $\text{SO}_3$  is produced from the reaction between 31.5-g of  $\text{S}_8$  and 8.63-g of  $\text{O}_2$ ?

12. What mass of  $\text{SO}_3$  is produced from the reaction of 12.4-g of  $\text{SO}_2$  and 3.45-g of  $\text{O}_2$ ?

13. What mass of  $\text{H}_2\text{SO}_4$  is produced from the reaction of 6.58-g of  $\text{SO}_3$  and 1.64-g of  $\text{H}_2\text{O}$ ?

14. What mass of  $\text{CdS}$  is produced if 8.47-g of cadmium reacts with 2.51-g of Sulfur?

# Prentice Hall Chemistry Chapter 11 Test Answers

**L Darling-Hammond**



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