

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## 17

## THERMOCHEMISTRY

**SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK (pages 505–510)**

*This section explains the relationship between energy and heat, and distinguishes between heat capacity and specific heat.*

**► Energy Transformations (page 505)**

- What area of study in chemistry is concerned with the heat transfers that occur during chemical reactions? thermochemistry
- Where the use of energy is concerned (in a scientific sense), when is work done?  
Work is done when a force is used to move an object.
- Circle the letter next to each sentence that is true about energy:
  - ☒ Energy is the capacity for doing work or supplying heat.
  - ☒ Energy is detected only because of its effects.
  - ☐ Heat is energy that transfers from one object to another because they are at the same temperature.
  - ☒ Gasoline contains a significant amount of chemical potential energy.
- Circle the letter next to each sentence that is true about heat.
  - ☒ One effect of adding heat to a substance is an increase in the temperature of that substance.
  - ☐ Heat always flows from a cooler object to a warmer object.
  - ☒ If two objects remain in contact, heat will flow from the warmer object to the cooler object until the temperature of both objects is the same.

**► Exothermic and Endothermic Processes (pages 506–507)**

- What can be considered the “system” and what are the “surroundings” when studying a mixture of chemicals undergoing a reaction? Write your answers where indicated below.

System: The mixture of chemicals itself is considered the system.

Surroundings: Everything but the mixture of chemicals is the surroundings, but practically speaking, the immediate vicinity of the system.

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**Daniel F McAuley**



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