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ATOMIC STRUCTURE

Chapter Test A

A. Matching

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

Column A	Column B
<u> J </u> 1. proton	a. the total number of protons and neutrons in the nucleus of an atom
<u> G </u> 2. atom	b. the weighted average mass of the atoms in a naturally occurring sample of an element
<u> A </u> 3. mass number	c. $\frac{1}{12}$ the mass of a carbon-12 atom
<u> C </u> 4. atomic mass unit	d. the number of protons in the nucleus of an element
<u> F </u> 5. electron	e. atoms with the same number of protons but different numbers of neutrons
<u> E </u> 6. isotopes	f. negatively charged subatomic particle
<u> D </u> 7. atomic number	g. the smallest particle of an element that retains its identity in a chemical reaction
<u> B </u> 8. atomic mass	h. a horizontal row of the periodic table
<u> H </u> 9. period	i. subatomic particle with no charge
<u> I </u> 10. neutron	j. positively charged subatomic particle

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. Which of the following is *not* a part of Dalton's atomic theory?
- All elements are composed of atoms.
 - Atoms of the same element are alike.
 - Atoms are always in motion.
 - Atoms that combine do so in simple whole-number ratios.
- _____ 12. The nucleus of an atom is
- negatively charged and has a low density.
 - negatively charged and has a high density.
 - positively charged and has a low density.
 - positively charged and has a high density.

Prentice Hall Chemistry Atomic Structure Test Answers

Robert R. Redfield



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