

## 14

## THE BEHAVIOR OF GASES

## Chapter Test II

## A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct description next to the item.

## Column I

1. The constant relating unit temperature change to the change in volume of a fixed amount of gas at constant pressure
2. The volume of a fixed amount of gas at constant pressure as the absolute temperature is changed
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used
4. An experimental gas constant determined from the ideal gas law at low pressures
5. A constant of nature that has values of about 8.314 J/mol·K
6. The a parameter of gas constant equations for determining the gas molecular weight of a gas
7. An experimental gas constant derived from the ideal gas law at low pressures
8.  $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. A constant of nature that has a value of 2.303
10. A constant of nature that has a value of about 8.314 J/mol·K

## Column II

- a. pressure gradient
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. universal gas constant, pressure
- h. volume
- i. compressibility
- j. distribution coefficient

## B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Subliming the solvent of a solid-solvent system gives a solid leaving a liquid solvent. Sublimation is used to separate a. the solid-solvent mixture b. the solvent from the solid c. the solid from the solvent d. the solvent from the solid

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