

Chapter 12 worksheet 4 (ws12.4) Colligative Properties

A pure substance has well-defined physical properties such as freezing and boiling points. In contrast, **many properties of a solution depend on the concentration of solute**. Properties of solutions that depend on the concentration (**but not on the identity**) of solute are called **colligative** (collective) properties. You will learn how the concentration of a solute affects four properties of a solution. As the concentration of solute increases:

vapor pressure decreases
boiling point increases
freezing point decreases
osmotic pressure increases

We will treat these effects both qualitatively and quantitatively and you will learn why they happen. We will discuss some important applications of these effects.

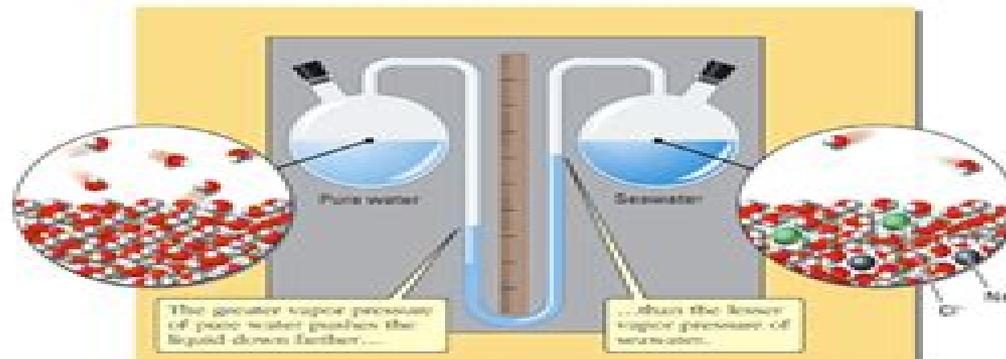
Volatile vs. non-volatile components of a solution

For each colligative property we will usually consider the case of a **non-volatile solute** dissolved in a **volatile solvent** (a typical example is **any ionic compound** dissolved in water).

A volatile substance is one with a measurable vapor pressure at moderate temperatures. In general, volatility increases as intermolecular forces decrease.

Vapor pressure reduction (Raoult's law)

Vapor pressure is the pressure inside a closed container when a liquid is in equilibrium with its own gas phase so the rate of evaporation is equal to the rate of condensation. When a **non-volatile solute** is dissolved in a **volatile solvent**, the vapor pressure of the solvent is reduced. As the concentration of the solute increases, the vapor pressure of the solvent decreases.



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Table of Contents Prentice Hall Chem Chapter 12 4 Answers

1. Understanding the eBook Prentice Hall Chem Chapter 12 4 Answers
 - The Rise of Digital Reading Prentice Hall Chem Chapter 12 4 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chem Chapter 12 4 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chem Chapter 12 4 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chem Chapter 12 4 Answers
 - Personalized Recommendations
 - Prentice Hall Chem Chapter 12 4 Answers User Reviews and Ratings
 - Prentice Hall Chem Chapter 12 4 Answers and Bestseller Lists
5. Accessing Prentice Hall Chem Chapter 12 4 Answers Free and Paid eBooks
 - Prentice Hall Chem Chapter 12 4 Answers Public Domain eBooks
 - Prentice Hall Chem Chapter 12 4 Answers eBook Subscription Services
 - Prentice Hall Chem Chapter 12 4 Answers Budget-Friendly Options

6. Navigating Prentice Hall Chem Chapter 12 4 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chem Chapter 12 4 Answers Compatibility with Devices
 - Prentice Hall Chem Chapter 12 4 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chem Chapter 12 4 Answers
 - Highlighting and Note-Taking Prentice Hall Chem Chapter 12 4 Answers
 - Interactive Elements Prentice Hall Chem Chapter 12 4 Answers
8. Staying Engaged with Prentice Hall Chem Chapter 12 4 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chem Chapter 12 4 Answers
9. Balancing eBooks and Physical Books Prentice Hall Chem Chapter 12 4 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chem Chapter 12 4 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chem Chapter 12 4 Answers
 - Setting Reading Goals Prentice Hall Chem Chapter 12 4 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chem Chapter 12 4 Answers
 - Fact-Checking eBook Content of Prentice Hall Chem Chapter 12 4 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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