

Name: Key

Class: _____

Date: _____

Stoichiometry Problem Set A

Directions: Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

Part 1: Straight Stoichiometry Problems

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

- 1.) For the reaction $2\text{HNO}_3 + 1\text{Mg(OH)}_2 \rightarrow 1\text{Mg(NO}_3)_2 + 2\text{H}_2\text{O}$, how many grams of magnesium ^B nitrate are produced from 8.00 mol of nitric acid (HNO_3)? A

$$\text{MM} = 148.3\text{g}$$

$$\frac{8.00\text{mol HNO}_3}{1} \times \frac{1\text{mol Mg(NO}_3)_2}{2\text{mol HNO}_3} \times \frac{148.3\text{g Mg(NO}_3)_2}{1\text{mol Mg(NO}_3)_2} = 592.8\text{g Mg(NO}_3)_2$$

- 2.) For the reaction $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$, how many moles of potassium ^B chlorate are required to produce 160. g of oxygen? A

$$\text{MM} = 32.0\text{g}$$

$$\frac{160.\text{g O}_2}{1} \times \frac{1\text{mol O}_2}{32.0\text{g O}_2} \times \frac{2\text{mol KClO}_3}{3\text{mol O}_2} = 3.33\text{ mole KClO}_3$$

- 3.) For the reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, how many grams of water are produced from 5.22 liters of oxygen gas? ^B A

$$\text{MM} = 18.0\text{g}$$

$$\frac{5.22\text{L O}_2}{1} \times \frac{1\text{mol O}_2}{22.4\text{L O}_2} \times \frac{2\text{mol H}_2\text{O}}{1\text{mol O}_2} \times \frac{18.0\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 8.39\text{g H}_2\text{O}$$

- 4.) In the reaction $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$, how many grams of hydrogen ^B are produced if 1.33 grams of sodium react completely? A

$$\text{MM} = 23.0\text{g}$$

$$\text{MM} = 2.0\text{g}$$

$$\frac{1.33\text{g Na}}{1} \times \frac{1\text{mol Na}}{23.0\text{g Na}} \times \frac{1\text{mole H}_2}{2\text{mole Na}} \times \frac{2.0\text{g H}_2}{1\text{mole H}_2} = 0.0578\text{g H}_2$$

- 5.) In the reaction $1\text{CO}_2 + 1\text{H}_2\text{O} \rightarrow 1\text{H}_2\text{CO}_3$, how many moles of carbonic ^B acid can be produced from 97.2 g of CO_2 ? A

$$\text{MM} = 44.0\text{g}$$

$$\frac{97.2\text{g CO}_2}{1} \times \frac{1\text{mole CO}_2}{44.0\text{g CO}_2} \times \frac{1\text{mole H}_2\text{CO}_3}{1\text{mole CO}_2} = 2.21\text{ moles H}_2\text{CO}_3$$

Did you achieve our target goal?

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

1
Feeling VERY confused

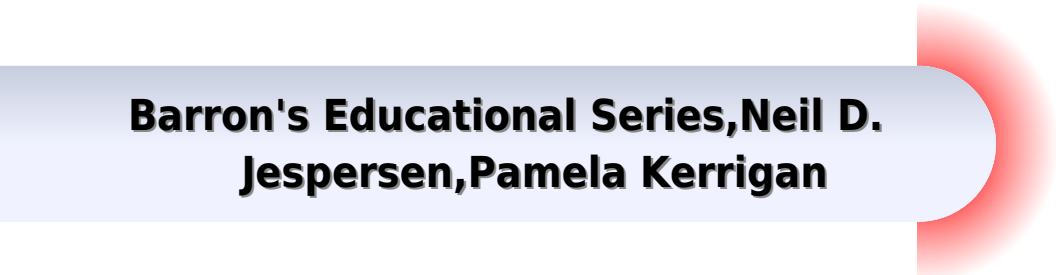
2
Not sure that I get this

3
I think I am catching on

4
Doing great! Bring it on!

Sample Problem Set Stoichiometry Chemistry

**Barron's Educational Series, Neil D.
Jespersen, Pamela Kerrigan**



Sample Problem Set Stoichiometry Chemistry:

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Chemistry in the Community (Enhanced Core Four) American Chemical Society, 2006-02-15

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Barron's Chemistry Practice Plus: 400+ Online Questions and Quick Study Review Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2022-07-05 Need quick review and practice to help you excel in Chemistry Barron's Chemistry Practice Plus features more than 400 online practice questions and a concise review guide that covers the basics of Chemistry Inside you'll find Concise review on the basics of Chemistry an excellent resource for students who want a quick review of the most important topics Access to 400 online questions arranged by topic for customized practice Online practice includes answer explanations with expert advice for all questions plus scoring to track your progress This essential guide is the perfect practice supplement for students and teachers

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18 at that time the disciples came to jesus saying who is the greatest in the kingdom of heaven 2 jesus called a little child to him and set him in their midst 3 and said truly i say to you unless you are converted and become like little children you will not enter the kingdom of heaven 4 therefore whoever humbles himself like

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18 then the lord appeared to him by a the terebinth trees of mamre as he was sitting in the tent door in the heat of the day 2 so he lifted his eyes and looked and behold three men were standing by him and when he saw them he ran from the tent door to meet them and bowed himself to the ground 3 and said my lord if i have now

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beginning with the warning that greatness in the kingdom of heaven is measured not by rank or power but by childlikeness mt 18 1 5 it deals with the care that the disciples must take not to cause the little ones to sin or to neglect them if they stray from the community mt 18 6 14 the correction of members who sin mt 18 15 18 the

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context of church discipline the context of matthew 18 18 20 is church discipline not prayer while it is true that the greek verb translated ask is often used in the context of prayer requests in the new testament jesus is not talking about prayer in this context

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aug 4 2017 in one particular manual binding and loosing prayers are encouraged on a daily basis and are applied not only in the face of alleged spiritual attack but also for a host of situations and conditions ranging from financial hardship tormenting thoughts and sexual sin to protection before anesthesia blood transfusion or surgery 1 while variou

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18 truly i tell you whatever you bind on earth will be a bound in heaven and whatever you loose on earth will be b loosed in heaven read full chapter footnotes matthew 18 18 or will have been matthew 18 18 in all english translations matthew

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18 receive this truth whatever you forbid on earth will be considered to be forbidden in heaven and whatever you release on earth will be considered to be released in heaven 19 again i give you an eternal truth if two of you agree to ask god for something in a symphony of prayer my heavenly father will do it for you 20 for wherever two o

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