

# REPORT SHEET

Name: \_\_\_\_\_  
 Course & Year: PSYCHOLOGY - 1A  
 Professor's Signature: [Signature]

Rating: 93%  
 Date: November 12, 2018  
 Group #: 2 - LECTURE

## Experiment No. \_\_\_\_\_ STOICHIOMETRY

*"Success is just a word unless you make it happen." -Dr. Naj*

- I. Objective: Determine the relationship between the amount of the reactants used and the products formed in a chemical reaction.
- II. Observation and Results:

Test tube no.	No. of drops of 1 M $Pb(NO_3)_2$	No. of drops of 1 M $Pb(NO_3)_2$	Height of ppt (mm)
1.	5 drops	5 drops	0.1
2.	10 drops	10 drops	0.2
3.	15 drops	15 drops	0.3
4.	20 drops	20 drops	0.4
5.	25 drops	25 drops	0.5

### QUESTIONS

1. Write the balanced equation for the reaction that took place. Indicate the gram-formula weight of the substances involved under their corresponding formulas in the equation.

$Pb(NO_3)_2 + Na_2CO_3 \rightarrow PbCO_3 + 2NaNO_3$   
 $Pb(NO_3)_2$   
 $Na_2CO_3$   
 $PbCO_3$   
 $2NaNO_3$

2. How do you explain the shape of the curve obtained when the amount of  $Pb(NO_3)_2$  solution was plotted against the height of the precipitate?

The shape of the curve obtained is linear, which shows that as the amount of  $Pb(NO_3)_2$  increases, the height of the precipitate also increases.

The graph shows a linear relationship between the height of the precipitate and the amount of  $Pb(NO_3)_2$  solution used, indicating a stoichiometric reaction.

*"What you focus on, expands." -Jim Rohn*

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**Judd E. Hollander**



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