

## SALT ANALYSIS



### SCHEME OF ANALYSIS

| EXPERIMENT  | OBSERVATION  | INFERENCE   |
|---|--|---|
| <b>PRELIMINARY EXAMINATIONS :</b>   |  |   |
| 1) Noted the colour   | Colourless   | Absence of $\text{Cu}^{2+}, \text{Fe}^{2+}, \text{Fe}^{3+}, \text{Mn}^{2+}, \text{Co}^{2+}, \text{Ni}^{2+}$   |
|   | Blue   | Pre. of $\text{Cu}^{2+}$  |
|   | Brown  | Pre. of $\text{Fe}^{3+}$  |
|   | Green  | Pre. of $\text{Ni}^{2+}$  |
|   | Pale pink  | Pre. of $\text{Mn}^{2+}$  |
|   | Pink   | Pre. of $\text{Co}^{2+}$  |
| 2) Noted the solubility<br>a) In water<br>b) In dil.HCl   | a) Soluble and no ppt with $\text{Na}_2\text{CO}_3$<br>b) Insoluble  | a) Pre. of $\text{NH}_4^+$<br>b) Pre. of $\text{Pb}^{2+}$   |
| 3) Flame Test :<br>A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame.            | a) Blue surrounded by green<br>b) Brick red<br>c) Crimson red<br>d) Bluish white<br>e) Grassy green on prolonged heating | a) Pre. of $\text{Cu}^{2+}$<br>b) Pre. of $\text{Ca}^{2+}$<br>c) Pre. of $\text{Sr}^{2+}$<br>d) Pre. of $\text{Pb}^{2+}$<br>e) Pre. of $\text{Ba}^{2+}$ |
| 4) Ash Test :<br>A little of the salt is heated with con. $\text{HNO}_3$ and cobalt nitrate solution. A filter paper is dipped in it and burnt. | a) Green ash<br>b) Blue tinted ash<br>c) Pale pink ash   | a) Pre. of $\text{Zn}^{2+}$<br>b) Pre. of $\text{Al}^{3+}$<br>c) Pre. of $\text{Mg}^{2+}$   |
| <b>SYSTEMATIC ANALYSIS OF ANIONS :</b>  |  |   |
| 1) $\text{dil.H}_2\text{SO}_4$ test :<br>To the salt add $\text{dil.H}_2\text{SO}_4$ , warm if required.  | a) Evolution of Colourless, odourless gas with brisk effervescence, turns limewater milky.<br>b) Smell of vinegar        | a) Pre. of carbonates<br>b) Pre. of acetate   |

# Salt Analysis Manual For P G Course

**Ying Liu**



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